

ORIGINAL RESEARCH PAPER

Predictions of the adsorption equilibrium of CO₂/O₂ mixture on multi-walled carbon nanotube using ideal adsorbed solution theory

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Abstract

Multiwalled carbon nanotubes (MWCNT) were found to be an effective separation media for purifying CO₂ from O₂. Significant uptakes of CO₂ and O₂ were measured at 288 K, 298K and 308 K over the pressure range of 1 to 40 bar using volumetric method in dual sorption vessels. The same shape of isotherms introduced a common mechanism of adsorption but the amount of CO₂ adsorbed on MWCNT is 2 times higher than O₂ adsorption. The mass uptake of CO₂ and O₂ by MWCNT was found to increase with increasing pressure and decreasing temperature. The experimental data was well fitted by the Langmuir and Freundlich model isotherms considering the values of regression correlation coefficients. Following a simple acidic treatment procedure, CO₂ and O₂ adsorption was increased over range of pressure. The adsorbents was characterized by N₂ adsorption isotherm at 77 K, Fourier transform infrared spectroscopy (FTIR), transmission electron microscopy (TEM), scanning electron microscopy (SEM) and thermogravimetric analysis (TGA). The effect of temperature and pressure on selectivity obtained from IAST demonstrated that maximum selectivity over the pressure and temperature ranges p = 0.5-5 bar and T = 298–308 K was achieved at 308 K and 5 bar.

Keywords: Adsorption isotherms; Carbon dioxide; IAST; MWCNT; Oxygen.

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INTRODUCTION

Human activities, such as the ignition of fuels have increased CO₂ concentrations in the atmosphere. CO₂ is one of the major greenhouse gases responsible for global warming [1]. Several technologies have been proposed to capture CO₂ including chemical absorption, cryogenic separation, physical adsorption, and membrane separation. Among these techniques, chemical absorption with aqueous solution is the most famous one [2, 3]. But disadvantages of this method cause physical adsorption plays an important role in the adsorption of CO₂. In the

design of an adsorption based separation process, the choice of the adsorbent is very important design consideration. Pressure swing adsorption (PSA) is a very promising separation and recovery process for this type of application [4, 5].

When the carbon nanotube (CNT) was discovered by Iijima in 1991 [6], because of their unique properties such as uniform porosity, high effective surface area and low mass density many researches were done on this new kind of porous adsorbents [7]. The internal and external adsorption sites of the curved nanotubes increase surface for adsorption and therefore influence the

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extent of adsorption [8].

The adsorption of gases inside microporous solid sorbents is the basis of many industrial processes for gas purification and gas separation. Although multicomponent adsorption can be determined experimentally, performing experiments that systematically explore the full range of potential operating conditions for a material of interest is time consuming. As a result, many works have been done on methods to predict the adsorption of mixtures of components using only pure component isotherm fits as data inputs [9]. Among these methods, the ideal adsorbed solution theory (IAST) of Myers and Prausnitz [10] is widely used to predict the multicomponent adsorption equilibria in a wide variety of systems.

In this study the potential of MWCNT as an adsorbent for CO₂ removal from flue gas (mixture of CO₂ and O₂) was studied. Also experimental results of CO₂ and O₂ adsorption by MWCNT at equilibrium state were described by Langmuir and Freundlich isotherms. Then, the ideal adsorbed solution theory in conjunction with Langmuir single-component isotherm model (Langmuir–IAS theory) was used to predict the adsorption of CO₂ and O₂ in the mixture and selectivity of CO₂/O₂ at three temperatures.

EXPERIMENTAL

Material and its characterization

MWCNT used in this study was purchased from Alpha Nanotechnologies Company, Ltd (China) which was synthesized with chemical vapor deposition (CVD) method with purity greater than 95%. The apparent density of MWCNT was 0.063 g.cm⁻¹. MWCNT was characterized with respect to wall thickness (single or multi-walled), diameter and length by TEM (Zeiss EM10C, Germany). Also, SEM (VEGA II LMU, Tescan, Czech Republic) was used to obtain supplementary information about the MWCNT before and after modification. The BET technique (Belsorp mini II model manufactured by BelJapan) employing N₂ adsorption isotherm at 77 K was used to measure textural properties of adsorbent. The surface functional groups of sorbents were characterized by FTIR (Bruker Vertex-70, Germany). Sample discs were prepared by mixing 1mg of the samples with 300 mg of KBr and scanned in the range of 400–4000 cm⁻¹. TGA (TGA-50, Shimadzu, Japan) was performed to study the thermal stability and dehydration characteristics of the adsorbents.

The scan temperature was maintained between 40 and 1000 °C at a heating rate of 10 °C. min⁻¹ in the atmosphere of air. Carbon dioxide, helium and O₂ with purity of 99.99%, 99.995% and 99.99% respectively, were purchased from Technical Gas Services, UAE.

ACID Modification

The modification of MWCNT was done by adding 100 mg of crude MWCNTs to 150 mL acid solution of H₂SO₄/HNO₃ (3:1 v/v) via sonication for 1 h. Then the solution was refluxed under stirring using magnetic stirrer equipped with reflux condenser, thermometer and oil bath for better heat distribution at 90 °C for 1 h. After cooling down to room temperature, the mixture was filtered and washed with distilled water several times until pH increased to neutral value (pH 7). The carboxylated MWCNT were then dried in a vacuum drying oven at 100 °C for 4 h.

Apparatus

The amount of CO₂ and O₂ adsorbed on the MWCNT were measured by using a volumetric method. Fig. 1 displays the apparatus that was designed to conduct the gas adsorption experiments. The apparatus was consisted of two high-pressure cells made of stainless steel, called as pressure and adsorption cell. They were placed in a thermostatic water bath (with working temperature range of 283-363K manufactured by Merck Company) to keep the temperature constant during the gas adsorption. The pressure cell was connected to a regulator and adsorption cell via a needle valve to control gas entrance. Gas flew from gas high-pressure reservoir to the pressure cell via the regulator by opening the connecting valve. The adsorption experiments were initiated by opening the valves between pressure cell and adsorption cell and allowing gas entrance to the adsorption cell loaded with the adsorbent. The pressure in the system was started to drop as gas was adsorbed on the adsorbent. The equilibrium state was reached when the pressure approached to a constant value. Both cells equipped with a PT100 temperature probe and a pressure transducer that were interfaced to the computer to monitor and record system pressure. The maximum allowable pressure in the installation was 50 bars and its working temperature was in the range of 283-343K.

Prior loading adsorbent into the adsorption

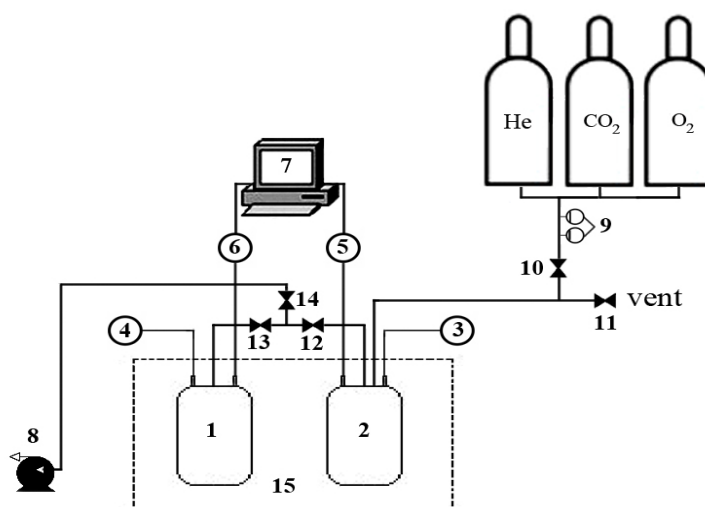


Fig.1. Schematic diagram of the volumetric adsorption set-up: (1) adsorption cell; (2) pressure cell; (3, 4) temperature digital indicator connected to the temperature probe; (5, 6) pressure digital indicator connected to the pressure transducer; (7) recorder; (8) vacuum pump; (9) Regulator; (10-14) valves.

cell, the setup was tested against any probable leak using pressurized air for 24 hours. Before each test, the adsorbent was degassed at 473K for 24 hours and the system was evacuated by vacuum pump to 0.1 mbar. Also the He gas was employed to determine the dead volume by using a calibration cell with a definite volume. By this method the exact volume of pressure and adsorption cell with their connection lines such as tubes and valves were measured.

The amount of adsorbed O₂ and CO₂ was calculated using material balance before and after each test and Soave-Redlich-Kwong (SRK) equation of state for the compressibility factor:

$$\left. \frac{PV}{ZRT} \right|_{L_1} + \left. \frac{PV}{ZRT} \right|_{a_1} = \left. \frac{PV}{ZRT} \right|_{L_2} + \left. \frac{PV}{ZRT} \right|_{a_2} + N \quad (1)$$

Where subscripts 1, 2, L and *a* stand for the initial state, final equilibrium state, pressure cell and adsorption cell, respectively. Also *V* is the volume, *P* is the pressure, *T* is the temperature, *R* is the universal gas constant, and *Z* is the gas compressibility factor.

RESULT AND DISCUSSION

Materials characterizations

The N₂ adsorption and desorption isotherms at 77K of the MWCNT is shown in Fig. 2. The isotherms were classified as the type II adsorption isotherm defined by IUPAC which is often encountered in adsorption by nonporous materials or on

materials with macropores or open voids [11]. At low partial pressure, the adsorbed volume has increased rapidly confirming the presence of a large number of micropores in this adsorbent. The main characteristics such as BET surface area, total pore volume and mean pore diameter of this adsorbent were achieved 294 m²/g, 0.6231 cm³/g and 4.68 nm respectively. The measurement of the pore size distribution using the Barrett-Joyner-Halenda (BJH) method provided the distribution of pores in the range 2–100 nm (Fig. 3). It indicated that MWCNT used in this study have both micro and meso pores.

It is obvious from Fig.4 (a) that the MWCNT had small amount of catalyst particles (as impurity). The caps of MWCNT were open; which are important for the gas adsorption because open tubes enable gas adsorption by both inside and outside of tubes [12]. The outer diameter of adsorbent ranges from 15 to 20 nm, and the inner diameter is about 4 nm. Its length was about 30 micrometer. The SEM image for MWCNT was shown in Fig. 4 (b). It is clearly demonstrated that via Van der Waals forces the MWCNT is held together into bundles and show high degree of aggregation. It was also due to CNT length as the MWCNT showed a larger degree of entanglement. From Fig. 4 (c), the morphology of MWCNT after modification remained undamaged and more obvious ends compared to the pristine MWCNT was observed. That is because oxidative etching along the walls of MWCNT by strong acids led to fragment of MWCNTs in lengths in the range

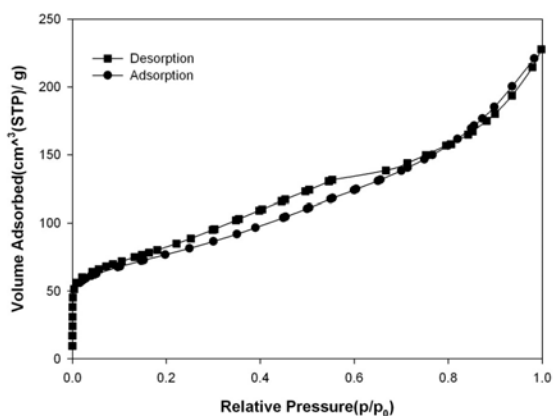


Fig. 2. N₂ adsorption/desorption isotherms of MWCNT.

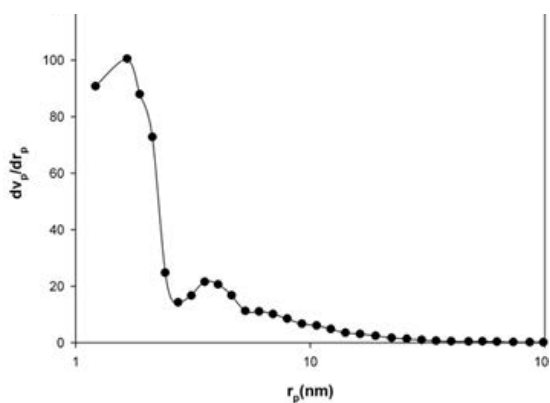


Fig. 3. Pore size distribution of the MWCNT using the BJH method.

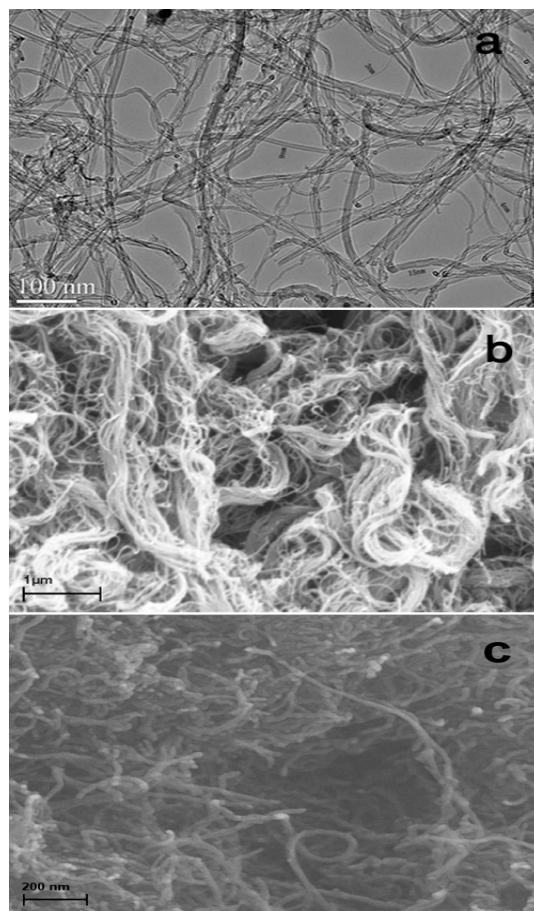


Fig. 4. TEM (a) and SEM images of MWCNT (b) and MWCNT-COOH (c).

of 100–300 nm.

The FTIR spectra of MWNT before and after the oxidation process is shown in Fig. 5. The peaks at 1635 cm⁻¹ is due to the C=C bond in aromatic rings of MWCNT. The peaks at 3440 cm⁻¹ is due to the OH band correspond to water molecules present in the KBr and unmodified MWCNT. After oxidation with strong acids, small peaks at 1714 cm⁻¹ appeared that belong to the C=O stretching vibration of COOH groups. The peaks at 2856 and 2925 cm⁻¹ are corresponded to the asymmetric and symmetric stretching vibrations of C-H. The C-O bond stretching can be found at 1122 cm⁻¹ [13-15].

The TGA curve of MWCNT and MWCNT-COOH is depicted in in Fig. 6. The figure shows a slight degradation below 600 °C, which is mainly due to evaporation of water molecules. Above 600 °C, the rate of degradation increases and corresponds to the beginning of decomposition of carbon

in MWCNT. The temperature at which onset of weight loss for the oxidized MWCNT occurred was lower than the temperature of unmodified MWCNT. That was due to the presence of the functional group bounded to the MWCNT surface [16].

Adsorption experiments

The adsorption studies were carried out at different pressures between 0 to 45 bar and 3 different temperatures, i.e. 288, 298 and 308 K. At any given temperature, the amount of gas adsorbed was a function of the pressure. Results demonstrated that gas uptake increased with an increase in pressure and a decrease in temperature. Fig. 7 illustrates the variation of O₂ and CO₂ equilibrium uptake by MWCNT adsorbent versus pressure at different temperatures. At higher pressures the adsorption isotherm curve of CO₂ on MWCNT increases almost linearly and

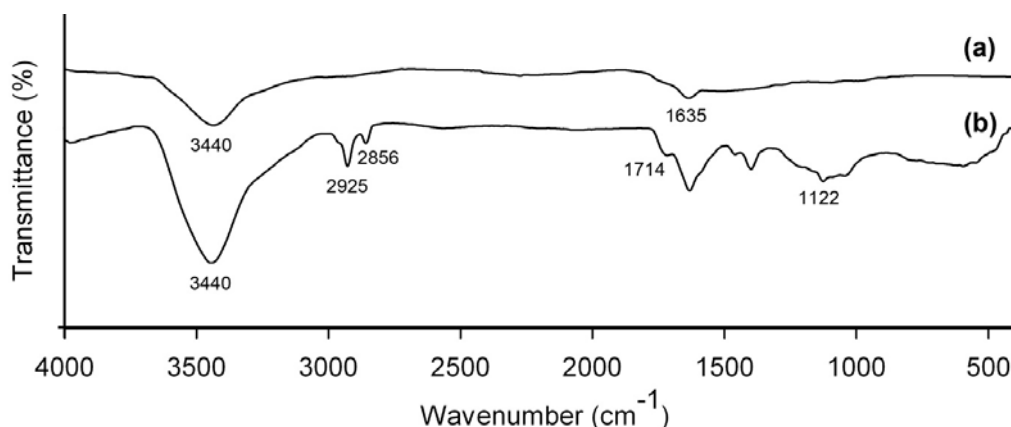


Fig. 5. FTIR spectra of MWCNT (a) and MWCNT-COOH (b).

it seems that the CO₂ uptake could be continued by increasing pressure, whereas the adsorption isotherm curve of O₂ on MWCNT leads to saturation at high pressure. The amount of CO₂ adsorption is more than O₂ adsorption because of its large quadrupole moment and high degrees of polarizability. In other words, the interactive forces between CO₂ and MWCNT are bigger than those, the interactive forces, for O₂. In this study the maximum amount of CO₂ and O₂ uptake by the MWCNT adsorbent was obtained 16.69 mmol.g⁻¹ and 8.65 mmol.g⁻¹ respectively at 288K and 50 bar. The results confirmed that by decreasing temperature to less than 288K and increasing pressure to above 50 bar, higher amounts of gas absorption can be achieved.

Adsorption isotherm models

Equilibrium data can be approximated using common and practical adsorption isotherms, which provide the basis for the design of adsorption systems. The amount of adsorbed material onto an adsorbent as a function of its pressure at constant temperature can be described by different adsorption isotherm models. The most famously used isotherm equation for modeling of the adsorption data is the Langmuir equation, which is based on monolayer sorption onto a surface with a finite number of equivalent sites and is given by the following equation [17]:

$$q_e = q_m \frac{K_L P}{1 + K_L P} \quad (2)$$

Where q_e is the amount of gas adsorbed per unit mass of MWCNT (mmol.g⁻¹); q_m is the maximum

amount of gas adsorbed (mmol.g⁻¹) and $K_L = k_a/k_b$, k_a and k_b are adsorption and desorption constants, respectively.

Freundlich adsorption isotherm is given by the following equation:

$$q_e = K_F P^{1/n} \quad (3)$$

Where K_F is Freundlich constant and $1/n$ is the heterogeneity factor. For $n > 1$, adsorption is favorable. The Freundlich isotherm indicates reversible adsorption process [18, 19].

Isotherm parameters were obtained through a nonlinear fit of experimental data with model isotherm equations (Fig. 7). Table 1 shows the values of Langmuir and Freundlich isotherms parameters. The comparison of the R² values from Table 1 shows that both Langmuir and Freundlich isotherm models yield better fit to the experimental data. With increasing the temperature, Langmuir parameters (q_m and K_L) decreased indicating that the affinity between both gases and adsorbent had an inverse relationship with temperature confirming an exothermic nature of the adsorption process. As the temperature increased from 288 to 308 K, the K_F value of the CO₂ and O₂ adsorbed by MWCNT decreased.

The effect of treatment

The effect of acid treatment on CO₂ and O₂ adsorption was studied at different pressures up to 40 bars and temperature of 298 K. Fig. 8 depicts the CO₂ and O₂ adsorption isotherm of acid-treated MWCNT. The experimental results indicate that the modified adsorbent shows better adsorption compared to pristine one. Acid modification of

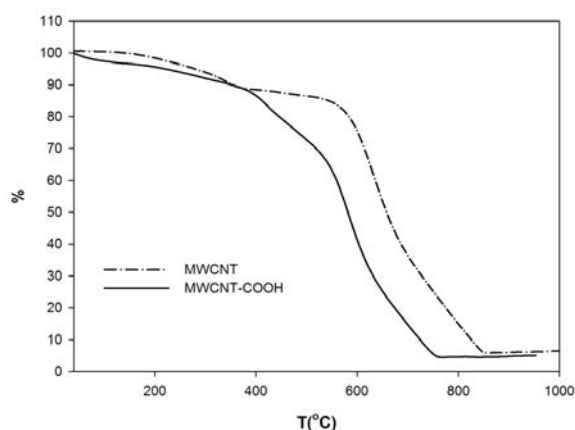


Fig. 6. TGA profile of MWCNT and MWCNT-COOH.

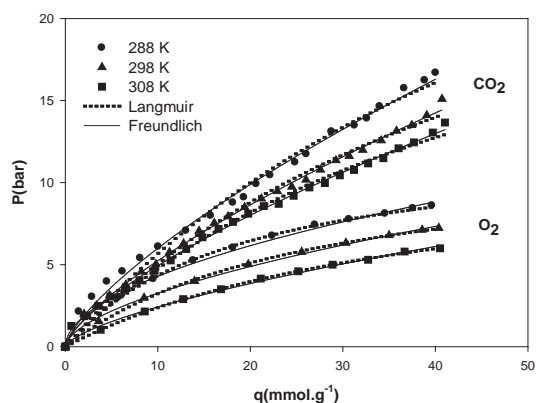


Fig. 7. CO₂ and O₂ adsorption isotherm for MWCNT at different temperatures and nonlinear fit of experimental data with Langmuir and Freundlich models.

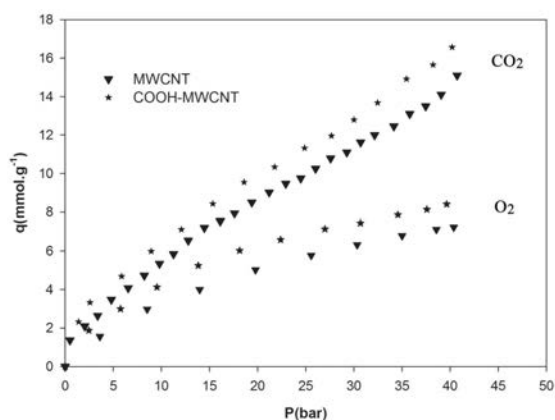


Fig. 8. The effect of acid modification in CO₂ and O₂ adsorption at 298 K.

adsorbent can eliminate effectively impurities such as the catalyst particles and amorphous carbon. Therefore the purity of CNT is increased and defects and cavities is formed which originate from removing impurities. Also, acid treatment opened closed-caps of MWCNT, which is useful for gas storage. Nevertheless, the treatments of MWCNT could not have great effects on the CO₂ and O₂ adsorption capacity (Fig.8) due to the covering the ends and sidewalls of MWCNT with various oxygen containing groups (mainly carboxyl groups) after pretreatment with concentrated acids. Similar trends in equilibrium adsorption of both gases on MWCNT and COOH-MWCNT verifies that the morphology of MWCNT remained undamaged after acid treatment.

The ideal adsorption solution theory (IAST)

As was mentioned before, the ideal adsorption solution theory (IAST) was the first major theory for predicting mixed gas adsorption from pure component isotherms. For the IAST application, the main condition is the availability of (i) single component adsorption isotherm data, and (ii) excellent curve fitting isotherm model for such data [20]. In this work, the Langmuir model was applied to fit the pure gas isotherms. The most important equations used in the IAST calculation are listed hereafter. The spreading pressure (π) is determined by:

$$\int_0^{P^{CO_2}} \frac{q_{CO_2}}{P} dp = \int_0^{P^{O_2}} \frac{q_{O_2}}{P} dp = \frac{\pi A}{RT} \quad (4)$$

where A is the surface area of the adsorbent The Raoult's law is given for both species:

$$P y_{CO_2} = P^0_{CO_2} x_{CO_2} \quad (5)$$

$$P y_{O_2} = P^0_{O_2} x_{O_2} \quad (6)$$

By solving these three equations, simultaneously, x_i , y_i was defined. With the values of $P^0_{CO_2}$, $P^0_{O_2}$, x_{CO_2} and x_{O_2} determined, the following equations are employed to calculate q_{CO_2} , q_{O_2} , q_t and adsorption selectivity S.

$$\frac{1}{qt} = \frac{x_{CO_2}}{q_{CO_2} (P^0_{CO_2})} + \frac{x_{O_2}}{q_{O_2} (P^0_{O_2})} \quad (7)$$

$$q_{CO_2} = q_t x_{CO_2} \quad (8)$$

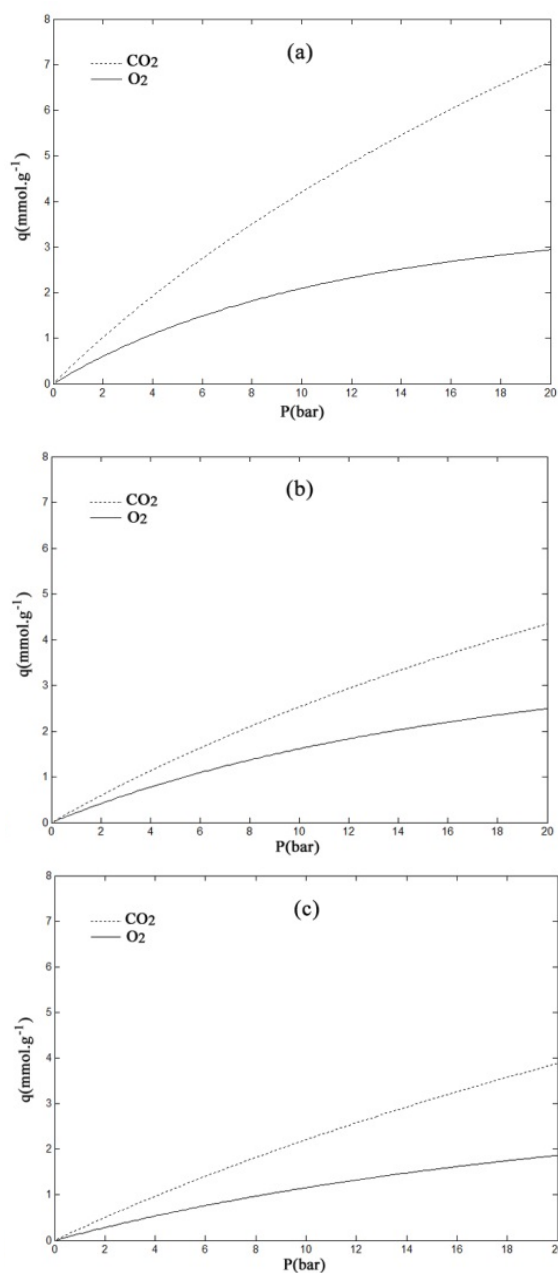


Fig.9. Binary adsorption isotherms for 50:50 gas-phase CO₂/O₂ mixtures predicted by IAST at different pressures and temperatures.

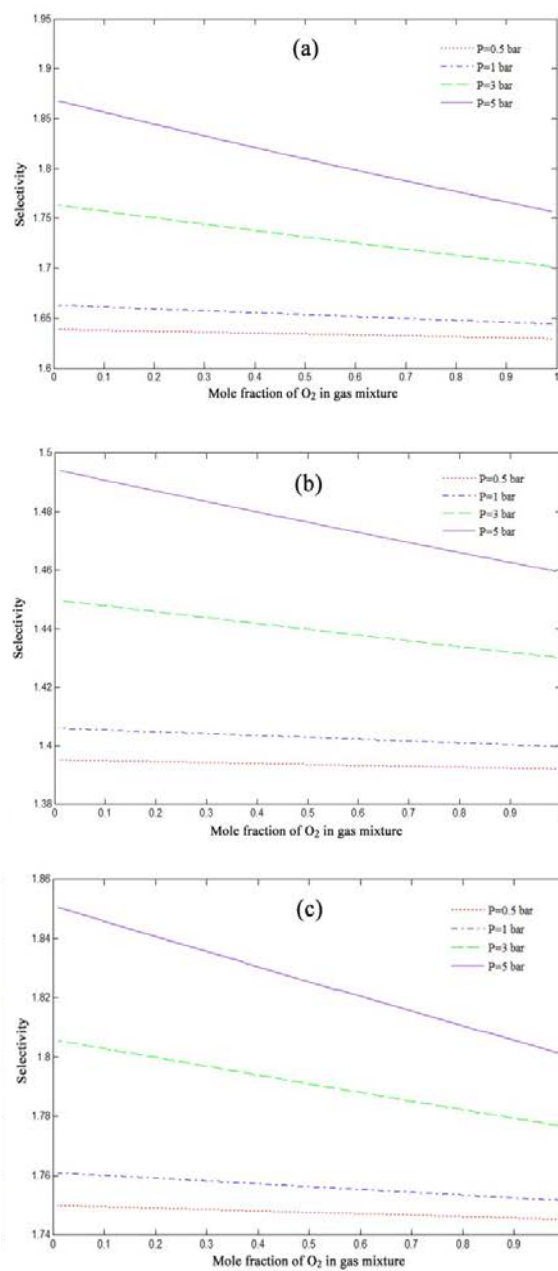


Fig. 10. Ideal selectivity of CO₂ to O₂ adsorbed on MWCNT at different temperature and different pressure for 50:50 gas-phase.

Table 1. Langmuir and Freundlich isotherm constants for adsorption of CO₂ and O₂ on MWCNT

T(K)	CO ₂						O ₂					
	Langmuir			Freundlich			Langmuir			Freundlich		
	q _m	K _L	R ²	K _F	n	R ²	q _m	K _L	R ²	K _F	n	R ²
288	41.31	0.026	0.984	1.140	1.587	0.995	12.55	0.053	0.997	1.294	1.919	0.996
298	30.81	0.020	0.987	1.110	1.451	0.996	12.03	0.037	0.999	0.851	1.712	0.998
308	28.79	0.018	0.992	0.992	1.434	0.996	11.92	0.025	1	0.527	1.504	0.997

$$q_{O_2} = q_T x_{O_2} \quad (9)$$

$$S_{CO_2/O_2} = \frac{x_{CO_2} / y_{CO_2}}{x_{O_2} / y_{O_2}} \quad (10)$$

where x_{CO_2} and x_{O_2} are the mole fractions in adsorbed phase and y_{CO_2} and y_{O_2} are the mole fractions of CO₂ and O₂ in gas phase. The adsorption isotherms of CO₂/O₂ mixtures were predicted by IAST at three temperatures for 50% O₂ in the bulk phase and showed in Fig. 9.

From Fig. 9, it can be concluded that if a CO₂/O₂ mixture is put into contact with this MWCNT, CO₂ would be adsorbed preferentially leading to an enrichment of O₂ in the gas phase. The CO₂/O₂ binary isotherms present behavior similar to the pure gas isotherms, with CO₂ as the dominant adsorbate. The predicted CO₂ and O₂ binary isotherms decreased relative to the pure component isotherms.

The selectivities for three different temperatures were predicted for 50:50 gas-phase CO₂/O₂ mixtures using IAST and was shown in Fig. 10. In all temperature, with increase in pressure or composition of the CO₂, selectivity increases. At higher compositions of CO₂ (lower composition of O₂), due to the lateral interactions between the adsorbed molecules of CO₂, the selectivity increased [24]. With increasing the temperature from 288 K to 298 K the selectivity was reduced from 1.81 to 1.48 for 50:50 gas-phase and pressure of 5 bar. Whereas by increasing the temperature from 298 K to 308 K the selectivity was increased from 1.48 to 1.83 for same condition.

Actually, based on the two main mechanisms for adsorption selectivity which are kinetic separation and thermodynamic separation, CO₂ has smaller kinetic diameter, higher quadrupole moment and higher polarizability compared to O₂ which in many cases result in stronger interaction with pore surface of the adsorbent. These data indicate that MWCNT can be successfully applied in CO₂/O₂ separation.

CONCLUSION

Experimental isotherm of pure CO₂ and O₂ adsorption was determined by the measurement of equilibrium uptake of CO₂ and O₂ at pressure range of 0–40 bar and at different temperatures.

Increasing the temperature from 288 to 308 K, decreased the gas capacity for both components due to the exothermic interaction. Acid treatment increased the amount of CO₂ and O₂ capture over entire range of pressure for MWCNT without any real damage and destruction on the morphology of MWCNT.

The CO₂/O₂ binary isotherms obtained from IAST present behavior similar to the pure gas isotherms, with CO₂ as the dominant adsorbate. According to the pure gas isotherm data on MWCNT, carbon dioxide is adsorbed much more than oxygen. The selectivity predicted from IAST indicates that MWCNT used in this study is selective to carbon dioxide. Both temperature and pressure are very important for selectivity. With increase of CO₂ concentration in feed gas, the selectivity of CO₂ increased. Maximum selectivity over the pressure and temperature ranges $p = 0.5$ – 5 bar and $T = 298$ – 308 K was achieved at 308 K and 5 bar. The results indicate that MWCNT can be successfully applied in CO₂/O₂ separation.

CONFLICT OF INTEREST

The authors declare that there are no conflicts of interest regarding the publication of this manuscript.

REFERENCES

- Samanta, A., A. Zhao, G.K. Shimizu, P. Sarkar, and R. Gupta, 2011. Post-combustion CO₂ capture using solid sorbents: a review. *Industrial & Engineering Chemistry Research*, 51(4): 1438-1463.
- Mondal, M.K., H.K. Balsora, and P. Varshney, 2012. Progress and trends in CO₂ capture/separation technologies: a review. *Energy*, 46(1): 431-441.
- Starr, K., X. Gabarrell, G. Villalba, L. Talens, and L. Lombardi, 2012. Life cycle assessment of biogas upgrading technologies. *Waste Management*, 32(5): 991-999.
- Sheikh, M.A., M.M. Hassan, and K.F. Loughlin, 1996. Adsorption equilibria and rate parameters for nitrogen and methane on Maxsorb activated carbon. *Gas separation & purification*, 10(3): 161-168.
- Grondein, A. and D. Bélanger, 2011. Chemical modification of carbon powders with aminophenyl and aryl-aliphatic amine groups by reduction of in situ generated diazonium cations: applicability of the grafted powder towards CO₂ capture. *Fuel*, 90(8): 2684-2693.
- Cao, D. and W. Wang, 2007. Storage of hydrogen in single-walled carbon nanotube bundles with optimized parameters: effect of external surfaces. *International Journal of Hydrogen Energy*, 32(12): 1939-1942.
- Reddy, A.L.M. and S. Ramaprabhu, 2008. Hydrogen adsorption properties of single-walled carbon nanotube-Nanocrystalline platinum composites. *International Journal of Hydrogen Energy*, 33(3): 1028-1034.
- Su, F., C. Lu, W. Cnen, H. Bai, and J.F. Hwang, 2009. Capture of CO₂ from flue gas via multiwalled carbon nanotubes.

- Science of the total environment, 407(8): 3017-3023.
9. Mofarahi, M. and F. Gholipour, 2014. Gas adsorption separation of CO₂/CH₄ system using zeolite 5A. Microporous and Mesoporous Materials, 200: 1-10.
 10. Myers, A. and J.M. Prausnitz, 1965. Thermodynamics of mixed-gas adsorption. AIChE Journal, 11(1): 121-127.
 11. Tsai, W., J. Yang, C. Lai, Y. Cheng, C. Lin, and C. Yeh, 2006. Characterization and adsorption properties of eggshells and eggshell membrane. Bioresource technology, 97(3): 488-493.
 12. Mishra, P., S. Mekala, F. Dreisbach, B. Mandal, and S. Gumma, 2012. Adsorption of CO₂, CH₄ and N₂ on a zinc based metal organic framework. Separation and Purification Technology, 94: 124-130.
 13. Vuković, G., A. Marinković, M. Obradović, V. Radmilović, M. Čolić, R. Aleksić, and P.S. Uskoković, 2009. Synthesis, characterization and cytotoxicity of surface amino-functionalized water-dispersible multi-walled carbon nanotubes. Applied Surface Science, 255(18): 8067-8075.
 14. Ma, P.C., J.-K. Kim, and B.Z. Tang, 2006. Functionalization of carbon nanotubes using a silane coupling agent. Carbon, 44(15): 3232-3238.
 15. Jin, S.H., Y.-B. Park, and K.H. Yoon, 2007. Rheological and mechanical properties of surface modified multi-walled carbon nanotube-filled PET composite. Composites Science and Technology, 67(15): 3434-3441.
 16. Murugesan, S., K. Myers, and V.R. Subramanian, 2011. Amino-functionalized and acid treated multi-walled carbon nanotubes as supports for electrochemical oxidation of formic acid. Applied Catalysis B: Environmental, 103(3): 266-274.
 17. Chowdhury, S., R. Mishra, P. Saha, and P. Kushwaha, 2011. Adsorption thermodynamics, kinetics and isosteric heat of adsorption of malachite green onto chemically modified rice husk. Desalination, 265(1): 159-168.
 18. Saygılı, H., F. Güzel, and Y. Önal, 2015. Conversion of grape industrial processing waste to activated carbon sorbent and its performance in cationic and anionic dyes adsorption. Journal of Cleaner Production, 93: 84-93.
 19. Ong, S.-A., L.-N. Ho, Y.-S. Wong, and A. Zainuddin, 2013. Adsorption behavior of cationic and anionic dyes onto acid treated coconut coir. Separation Science and Technology, 48(14): 2125-2131.
 20. Chen, H. and D.S. Sholl, 2007. Examining the accuracy of ideal adsorbed solution theory without curve-fitting using transition matrix Monte Carlo simulations. Langmuir, 23(11): 6431-6437.