

Experimental and Theoretical Aspects of Azo Dye Degradation by UV/H₂O₂ Process: Review and experimental comparison of some kinetic rate expressions

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Abstract

Due to their unique properties, azo dyes are extensively used, especially in the textile industry. Due to their low biodegradability, these compounds cannot be treated in wastewater treatment plants and discharges of these effluents poses a serious threat to the receiving water bodies. In the literature, several advanced oxidation processes have been studied for decolorization and mineralization of these toxic compounds. Among these advanced oxidation processes, the UV/H₂O₂ process has attracted great attention with its high efficiency in removing these compounds. The goal of this paper is to review the kinetic rate expressions developed to describe azo dye degradation by UV/H₂O₂ process. A detailed review of pseudo-first-order reaction mechanism, as well as reactor design models, is provided. Finally, a set of experiments are conducted with Reactive Black 5 to compare the model estimations with the observed data. In addition, a regression model is developed using response surface methodology to optimize operating conditions. The experimental results indicate that the optimum pH value that gives the maximum reaction constant is 5.74. Moreover, initial dye concentration is found to be a more significant parameter for decay rate constant than pH value. The open questions and future research topics are also discussed.

Keywords: Azo dye, Photolysis, •OH, Response surface methodology, UV/H₂O₂

Introduction

Azo dyes are the largest group of synthetic dyes, accounting for more than 70% percent of all the commercial dyes produced worldwide (Chung, 2016). In 1991, the Food and Drug Administration has certified more than 3000 tons of azo dyes for use in foods, pharmaceuticals and personal care products. There are currently over 2000 azo dyes and more than 7×10^5 tons of these dyes are manufactured globally and extensively used in the textile industries (Iark et al., 2019). A certain amount of them is inevitably lost in the process of their manufacturing and left in environmentally hazardous industrial wastes. The main source of dye water pollution in water is the unfixed excess dyes that are discharged with the significant volumes of water as textile effluents (Körbahti, 2007; Hassaan and Nemr, 2017).

The chemical structures of the azo dyes are distinguished by highly substituted aromatic rings joined by one or more azo groups which makes them recalcitrant to the conventional wastewater treatment processes (Song et al., 2010). Hence, they are discharged with the wastewater treatment plant effluents and end up in the receiving water bodies. Once these compounds are released into the water sources they disturb aquatic life by reducing the reoxygenation capacity of water and blocking sunlight. Azo dyes also have a significant environmental effect due to their degradation products, including those known to be highly carcinogenic

aromatic amines (Chiu et al., 2019). Recent studies show that these dyes have harmful impacts on the environment as well as on human health (Hassaan and Nemr, 2017). The removal of azo dyes from textile wastewater before discharge is thus a serious environmental concern. Common approaches such as chemical coagulation, air flotation, and adsorption can be used to remove the dyes from wastewater (Chafi et al., 2011; Lu et al., 2010). Such conventional approaches, however, primarily transfer the pollutants from one phase to another, without reducing their toxicity. Advanced oxidation processes on the other hand are alternative methods to completely remove and mineralize these recalcitrant pollutants (Cuerda-Correa et al., 2019).

Advanced Oxidation Processes (AOPs), which involves the generation of highly oxidative species, especially the hydroxyl radical (•OH) for the oxidation of a broad range of recalcitrant contaminants, has drawn the attention of environmental engineers in recent years (Muniyasamy et al., 2020). The most common AOPs for azo dye degradation are ozonation, UV/H₂O₂, sonication, Fenton processes and their combinations (Viswanathan, 2018). Among various combined processes the UV/H₂O₂ process has several advantages such as low capital cost and no solid waste generation (El-Dein et al., 2001). In this process, photolysis of the added H₂O₂ results in the formation of •OH which is recognized with non-selectivity and strong oxidizing capacity stronger than all conventional oxidizing agents.

Many authors have investigated the kinetics of degradation of azo dyes by UV/H₂O₂ and developed models for the interpretation of kinetic results (Muruganandham and Swaminathan, 2004; Gultekin and Ince, 2004; Shu et al., 2009; Aleboyyeh et al., 2003; Galindo and Kalt, 1999; El-Dein et al., 2001; Behnajady et al., 2004; Behnajady et al., 2006; Malik and Sanyal, 2004).

Response surface methodology (RSM) is considered as the most effective for modeling, analysis and optimization of multivariable systems. A response surface model provides an approximation for the system response (e.g., percent yield of a reaction) for given values of each variable that are included in the study and has the potential to influence the response (e.g., temperature and pressure). Although a response surface model can include higher-order terms of independent variables, a full second-order polynomial empirical

model is widely used to describe multifactor chemical processes (Hartley 1959; Deming 1988).

This paper aims to review the kinetic rate expressions developed for homogeneous photocatalytic azo dye degradation by the UV/H₂O₂ process. The kinetic rate expressions reviewed in the first section of the paper are given in Table 1. Furthermore, some of the expressions are tested using experimental data achieved through laboratory experiments and a Response Surface Model (RSM) design is used to determine the optimum pH that maximizes mineralization and decolorization. In addition, sensitivity analysis is performed using a linear regression model with Standardized Regression Coefficients (SRC). It is important to note that, reaction rate constant expressions of the reviewed studies are not linked to each other. In other words, the rate constant might be expressed with the same term (such as k_1 and k_2) but does not necessarily have the same meaning in the different studies reviewed.

Table 1. Some kinetic rate expressions for oxidation of azo dyes with UV/H₂O₂.

Reference	Expression
Colonna et al. (1999)	$-\frac{d[C_{dye}]}{dt} = \frac{2 \cdot \Phi \cdot k_1 \cdot I_a \cdot [C_{dye}]}{(k_1 - k_2) \cdot [C_{dye}] + k_2 \cdot [C_{dye}]_0}$
Behnajady et al. (2004)	$-\frac{d[C_{dye}]}{dt} = \frac{2 \cdot \Phi \cdot k_1 \cdot I_0 \cdot F_{H2O2}}{k_2 \cdot [C_{dye}]_0} \cdot [C_{dye}]$ $F_{H2O2} = \epsilon_{H2O2} \cdot [H_2O_2] / (\epsilon_{H2O2} \cdot [H_2O_2] + \epsilon_{dye} \cdot [C_{dye}])$
Malik and Sanyal, (2004)	$k_{obs} = 8.11[H_2O_2]_0^{0.63}$
El-Dein et al. (2001)	$-\frac{dC_{dye}}{dt} = \frac{k_{10} C_{H2O2}}{k_2 + C_{H2O2}} C_{dye} I_0$
Behnajady et al. (2006)	$-\frac{dC_{dye}}{dt} = (2.10^{-4} \cdot I_0^{0.75} + 3.10^{-4} \cdot I_0^{1.38} \cdot [H_2O_2]_0^{0.49}) \cdot \psi_{dye} \cdot [C_{dye}]$ $-\frac{dC_{dye}}{dt} = (2.10^{-4} \cdot I_0^{0.75} + 0.1 \cdot I_0^{1.38} \cdot [H_2O_2]_0^{-0.39}) \cdot \psi_{dye} \cdot [C_{dye}]$
Shen and Wang (2002)	$\int_{C_{dye,in}}^{C_{dye,eff}} \frac{dC_{dye}}{C_{dye}^p e^{-\epsilon_{dye} \cdot C_{dye} \cdot m \cdot (r-ri)}} = k [I_0 \cdot \frac{ri}{r} e^{-(E+\epsilon_{H2O2} \cdot C_{H2O2}) \cdot (r-ri)}]^m \cdot [H_2O_2]^n \cdot L / v_z$
Chang et al. (2010)	$-r_{dye0} = \frac{k_{01} k_a I_0 \exp(-k_H [H2O2] - k_D [Dye]) [H2O2] [Dye] (1 + k_b [H2O2])}{[H2O2] + k_a [Dye] + k_a k_b [Dye] [H2O2]}$

Current status of kinetic rate expressions for photocatalytic dye oxidation

Pseudo-First-Order Models

Several studies revealed that the azo dye degradation by UV/H₂O₂ can be approximated with a first-order decay (Muruganandham and Swaminathan, 2004; Gultekin and Ince, 2004; Shu et al., 2009; Neamtu et al., 2002; Aleboyyeh et al., 2003). Taking into account the effect of hydroxyl radical concentration the main kinetic pathway of azo dyes removal is expressed as follows:

$$-\frac{dC_{dye}}{dt} = k_2 \cdot C_{dye} \cdot C_{OH\cdot} \quad (1)$$

where the concentration of azo dye is expressed in C_{dye} and the concentration of $OH\cdot$ is denoted by $C_{OH\cdot}$ (mol/L).

The $C_{OH\cdot}$ can be considered to be a constant in the presence of excess hydrogen peroxide by pseudo-stationary hypothesis (Neamtu et al., 2002). Equation 1, therefore, is simplified into a kinetic model of the pseudo-first-order given in Equation 2:

$$-\frac{dC_{dye}}{dt} = k_1 \cdot C_{dye} \quad (2)$$

In Equation 2, k_1 is the pseudo-first-order rate constant. Pseudo steady-state approximation, in which the concentration of hydroxyl radicals ($OH\cdot$) does not change with the reaction time, is commonly used (Kumoro et al., 2017; Rajchel-Mieldzioć et al., 2020; Fard et al., 2016; El-Dein et al., 2003). This approach has been widely used to eliminate the nonmeasurable radicals concentration, correlating it as a function of H₂O₂ concentration (Neamtu et al, 2002; El-Dein et al., 2003).

Colonna et al. (1999) proposed a kinetic model taking into account that decolorization is complete in a relatively shorter amount of time compared to the duration that mineralization requires. According to that hypothesis, fluorescent intermediates are generated initially. Further reactions with hydroxyl radicals and photochemical processes then lead to smaller organic molecules, that will eventually be completely mineralized. The reactions are shown in Eq 3- 6:



In the study of Colonna et al. (1999), these chain reactions are expressed by a simplified model of which kinetic equations are given as follows:

$$-\frac{d[\text{H}_2\text{O}_2]}{dt} = 2 \cdot \phi \cdot I_a \quad (7)$$

$$-\frac{d[\text{C}_{\text{dye}}]}{dt} = k_1 \cdot [\text{C}_{\text{dye}}] \cdot [\text{C}_{\text{OH}\cdot}] \quad (8)$$

$$-\frac{d[\text{OH}\cdot]}{dt} = k_1 \cdot [\text{C}_{\text{dye}}] \cdot [\text{C}_{\text{OH}\cdot}] + k_2 \cdot [\text{Int}] \cdot [\text{C}_{\text{OH}\cdot}] - 2 \cdot \phi \cdot I_a \quad (9)$$

$$-\frac{d[\text{Int}]}{dt} = k_2 \cdot [\text{Int}] \cdot [\text{C}_{\text{OH}\cdot}] - k_1 \cdot [\text{C}_{\text{dye}}] \cdot [\text{C}_{\text{OH}\cdot}] \quad (10)$$

$$-\frac{d[\text{P}_1]}{dt} = k_2 \cdot [\text{Int}] \cdot [\text{C}_{\text{OH}\cdot}] + k_3 \cdot [\text{P}_1] \quad (11)$$

$$\frac{d[\text{P}_2]}{dt} = k_3 \cdot [\text{P}_1] \quad (12)$$

where ϕ is the quantum yield of the photochemical hydrogen peroxide dissociation, I_a is the H_2O_2 -absorbed radiation intensity.

In that study k_2 is the second-order kinetic constant for the hydroxyl radical attack to intermediate products. Equation 13 can be derived from the aforementioned kinetic equations by applying steady-state approximation to $[\text{OH}\cdot]$ (Colonna et al., 1999).

$$-\frac{d[\text{C}_{\text{dye}}]}{dt} = \frac{2 \cdot \phi \cdot k_1 \cdot I_a \cdot [\text{C}_{\text{dye}}]}{(k_1 - k_2) \cdot [\text{C}_{\text{dye}}] + k_2 \cdot [\text{C}_{\text{dye}}]_0} \quad (13)$$

where $[\text{C}_{\text{dye}}]_0$ is the initial dye concentration and $[\text{Int}] + [\text{C}_{\text{dye}}] \approx [\text{C}_{\text{dye}}]_0$.

k_2 has been reported to be of the order of $10^9 \text{ L mol}^{-1} \text{ s}^{-1}$ (Buxton et al., 1988; Kochany et al., 1992). Therefore neglecting $(k_1 - k_2)$ in the denominator of Equation 13 yields:

$$-\frac{d[\text{C}_{\text{dye}}]}{dt} = \frac{2 \cdot \phi \cdot k_1 \cdot I_a \cdot [\text{C}_{\text{dye}}]}{k_2 \cdot [\text{C}_{\text{dye}}]_0} = \frac{2 \cdot \phi \cdot k_1 \cdot I_a \cdot [\text{C}_{\text{dye}}]}{k_2 \cdot [\text{C}_{\text{dye}}]_0} \cdot F \cdot (1 - 10^{-A}) \quad (14)$$

where F is the fraction of light absorbed by H_2O_2 , A represents the total absorbance of the solution within the same wavelength range, I_a is the intensity of the radiation absorbed by the sample and I_0 is the radiation intensity impinging on the solution.

Due to the high absorption of the dye during the initial part of the process, the following can be taken for consideration under continuous irradiation ($1 - 10^{-A}$) ≈ 1 so Equation 14 becomes:

$$-\frac{d[\text{C}_{\text{dye}}]}{dt} = \frac{2 \cdot \phi \cdot k_1 \cdot I_0 \cdot F_{\text{H}_2\text{O}_2}}{k_2 \cdot [\text{C}_{\text{dye}}]_0} \cdot [\text{C}_{\text{dye}}] \quad (15)$$

In that equation fraction of light absorbed by hydrogen peroxide ($F_{\text{H}_2\text{O}_2}$) is a parameter. $F_{\text{H}_2\text{O}_2}$ was explained in a more detailed manner in the study of Behnajady et al. (2004). In that study, the authors used the same model for azo dye degradation process. The major contribution of their study is a detailed explanation of absorbed radiation intensity by sample (I_a) which is given in Equations 16-17:

$$I_a = I_0 \cdot F_{\text{H}_2\text{O}_2} \cdot [1 - \exp(-2.3L \cdot (\epsilon_{\text{H}_2\text{O}_2} \cdot [\text{H}_2\text{O}_2] + \epsilon_{\text{dye}} \cdot [\text{C}_{\text{dye}}]))] \quad (16)$$

$$F_{\text{H}_2\text{O}_2} = \frac{\epsilon_{\text{H}_2\text{O}_2} \cdot [\text{H}_2\text{O}_2]}{\epsilon_{\text{H}_2\text{O}_2} \cdot [\text{H}_2\text{O}_2] + \epsilon_{\text{dye}} \cdot [\text{C}_{\text{dye}}]} \quad (17)$$

where L is the optical path length of the system, $\epsilon_{\text{H}_2\text{O}_2}$ and ϵ_{AO7} are the molar extinction coefficients for H_2O_2 and the dye, respectively.

In their expression authors state that the exponential terms given in brackets in Equation 16 are negligible due to high absorbance of the dye and H_2O_2 in the initial part of the process. So it can be written as:

$$1 - \exp(-2.3L \cdot (\epsilon_{\text{H}_2\text{O}_2} \cdot [\text{H}_2\text{O}_2] + \epsilon_{\text{dye}} \cdot [\text{C}_{\text{dye}}])) \approx 1 \quad (18)$$

Substituting Equation 18 and 16 into Equation 15 yields Equation 19:

$$-\frac{d[\text{C}_{\text{dye}}]}{dt} = k_{\text{ap}} \cdot [\text{C}_{\text{dye}}] \quad (19)$$

where

$$k_{\text{ap}} = \frac{2 \cdot \phi \cdot k_1 \cdot I_0 \cdot F_{\text{H}_2\text{O}_2}}{k_2 \cdot [\text{C}_{\text{dye}}]_0} \quad (20)$$

Note that in these equations the pseudo-first-order rate constant (k_{ap}) is a function of H_2O_2 and dye concentration. On the other hand results of experiments in that study indicate that k_{ap} increases with an increasing quantity of H_2O_2 and reaches nearly a constant value above optimum concentration. Another study that aims to investigate the dependence of k on the concentration of H_2O_2 and the radiation flux was conducted by El-Dein et al. (2001). In agreement with Behnajady et al. (2004), experimental results showed that at low H_2O_2 concentrations, observed k increased proportionally to the H_2O_2 concentration. A maximum rate has been observed at higher concentrations. They described that nonlinear relationship with the following kinetic expression:

$$k = \frac{k_1 \cdot C_{\text{H}_2\text{O}_2}}{k_2 + C_{\text{H}_2\text{O}_2}} \quad (21)$$

where both k_1 and k_2 can depend on I_0 . For the case that $C_{\text{H}_2\text{O}_2} \gg k_2$ Equation 21 becomes $k = k_1$.

The authors plotted k_1 versus UV intensity and observed a linear relationship which can be expressed as:

$$k_1 = k_a \cdot I \quad (22)$$

Therefore Equation 22 was written as:

$$-\frac{d[\text{C}_{\text{dye}}]}{dt} = \frac{k_a \cdot C_{\text{H}_2\text{O}_2}}{k_2 + C_{\text{H}_2\text{O}_2}} \cdot C_{\text{dye}} \cdot I \quad (23)$$

In that study, El-Dein et al. (2001) developed a reaction rate expression taking into account the effect of both H_2O_2 and UV flux irradiation based on reaction kinetics but neglected the effect of hydrogen peroxide dissociation and initial dye concentration. However, several studies are showing the dependence of the reaction rate to the initial dye concentration (Galindo and Kalt, 1999; Kdasi et al., 2004; Shu et al., 1994). The effect of initial dye concentration will also be tested with the experimental data in the last section of this paper.

Malik and Sanyal, (2004) carried out a kinetic study considering the contribution of the UV light by itself and the combined UV light and H₂O₂ reaction. They considered that during UV/H₂O₂ decolorization of azo dye two reactions take place in parallel; a pure photolysis reaction and an H₂O₂-assisted oxidation reaction. Thus, the overall rate of decolorization was expressed as:

$$-\frac{dC_{dye}}{dt} = k_{UV} \cdot C_{dye}^a \cdot I^b + k_{UV/H_2O_2} C_{dye}^c \cdot C_{H_2O_2}^d \cdot I^e \quad (24)$$

where C_{dye} is the dye concentration (M), k_{UV} is the rate constant of dye removal with UV radiation alone, I is the UV radiation intensity (kW/m²), k_{UV/H₂O₂} is the rate constant of dye removal with UV radiation in the presence of H₂O₂, C_{H₂O₂} is the concentration of H₂O₂ (M), and a-e are reaction orders.

It was hypothesized that H₂O₂ acted as a catalyst in the UV/H₂O₂ process because the change in concentration of H₂O₂ during dye removal was nearly constant. With that consideration and since the light intensity was kept constant, the rate expression (24) was simplified as:

$$-\frac{dC_{dye}}{dt} = (k'_{UV} + k'_{UV/H_2O_2}) C_{dye}^n \quad (25)$$

where k'_{UV}=k_{UV}·I^b and k'_{UV/H₂O₂}=k_{UV/H₂O₂}·C_{H₂O₂}^d·I^e; n is the order of reaction.

The results of the experiment in their study indicated that the value of n was 1. The rate expression can, therefore, be considered to be pseudo-first-order in dye concentration like other studies in the literature. The values of k_{obs} = (k'_{UV} + k'_{UV/H₂O₂}) and k'_{UV/H₂O₂} for various H₂O₂ concentrations were determined from experimental data using a pseudo-first-order kinetic model. A straight line was given by a logarithmic plot of k_{obs} against the initial concentration of H₂O₂ (R = 0.989). Thus, the pseudo-first-order rate constant (k_{obs}) as a function of H₂O₂ concentration up to the 5.88 x 10⁻³ M is obtained as:

$$k_{obs} = 8.11[H_2O_2]_0^{0.63} \quad (26)$$

In that study, the authors neglected the hydrogen peroxide dissociation by assuming that hydrogen peroxide concentration will be constant during the reaction. However, as previously explained, the aim of advanced oxidation is to generate highly reactive hydroxyl radicals by dissociating hydrogen peroxide. That mechanism is shown previously by Equation 7 and explained in the study of Colonna et al., 1999. For that reasons, constant hydrogen peroxide concentration assumption is difficult to justify in the UV/H₂O₂ process. Unlike Malik and Sanyal (2004), Behnajady et al. (2006) subdivided the UV/H₂O₂ process into three parts: the reaction rate caused by H₂O₂ alone, by UV photolysis alone and by UV/H₂O₂. The kinetic equation of dye degradation was given as follows:

$$-\frac{dC_{dye}}{dt} = k_1 \cdot [H_2O_2]_0^a [C_{dye}] + k_2 \cdot I_0^b \cdot [C_{dye}] + k_3 \cdot I_0^m \cdot [H_2O_2]_0^n [C_{dye}] \quad (27)$$

where k₁, k₂ and k₃ are the rate constants with H₂O₂ alone, UV alone and UV/ H₂O₂, respectively. I₀ represents the light intensity whereas a, b, m and n are the reaction orders.

Their results showed that the removal of dye was negligible with H₂O₂ alone (k₁≈0). Therefore equation 27 becomes:

$$-\frac{dC_{dye}}{dt} = (k_2 \cdot I_0^b + k_3 \cdot I_0^m \cdot [H_2O_2]_0^n) \cdot [C_{dye}] \quad (28)$$

The effect of initial dye concentration is inserted as a multiplicative effect on that equation. To consider this effect, a regression model in Equation 24 is fitted to a relative increase in k_{ap} for different initial dye concentrations where ψ_{dye} is named initial dye concentration index. In that regression model, dependent variables are the relative increase in k_{ap} values with respect to k_{ap} at 30 mg dye /L. Using least square method coefficients of the linear regression model, given in Equation 29, are calculated.

$$\psi_{dye} = 11.065 \cdot \frac{1}{[C_{dye}]_0} + 0.6532 \quad (29)$$

In both processes, a linear regression method using k_{ap} and ψ_{dye} values at low and high concentrations was employed to obtain reaction orders of H₂O₂ concentration and light intensity. With substituting the reaction orders for H₂O₂, light intensity and ψ_{dye} in the above reaction, rate expressions given in Equation 30-31 are obtained for degradation of dye in the UV/H₂O₂ process for 0 mg/L < [H₂O₂]₀ < 650 mg/L, and 650 mg/L < [H₂O₂]₀ < 1500 mg/L respectively. The light intensity in these equations varies from 0 to 44 W m⁻² while the initial dye concentration is between 10 and 30 mg/L.

$$-\frac{dC_{dye}}{dt} = (2.10^{-4} \cdot I_0^{0.75} + 3.10^{-4} \cdot I_0^{1.38} \cdot [H_2O_2]_0^{0.49}) \cdot \psi_{dye} \cdot [C_{dye}] \quad (30)$$

$$-\frac{dC_{dye}}{dt} = (2.10^{-4} \cdot I_0^{0.75} + 0.1 \cdot I_0^{1.38} \cdot [H_2O_2]_0^{-0.39}) \cdot \psi_{dye} \cdot [C_{dye}] \quad (31)$$

Hence, in the study of Behnajady et al. (2006) a rate equation was developed that correlates the reaction rate to light intensity, initial dye and H₂O₂ concentrations. Although the rate expression was empirical, the authors clarified that it provides detailed information about the influence of operational parameters on the reaction rate. Note that the H₂O₂ reaction order was reported as 0.69 for a mixture of three azo dyes in the study of Malik and Sanyal (2004). On the other hand, it is reported as -0.39 and 0.49 in that study for high and low initial H₂O₂ concentrations, respectively. The authors attributed this difference to different chemical structures of different azo dyes. However, the results can also be attributed to the difference in the range of H₂O₂ concentrations and UV light intensity, the use of different photocatalytic reactors with different designs and different operating conditions.

In the study of Chang et al. (2010), the effects of initial H_2O_2 and dye concentrations, pH, and power of UV radiation on decolorization rate were predicted by a kinetic model based on the reaction network analysis shown in Figure 1.

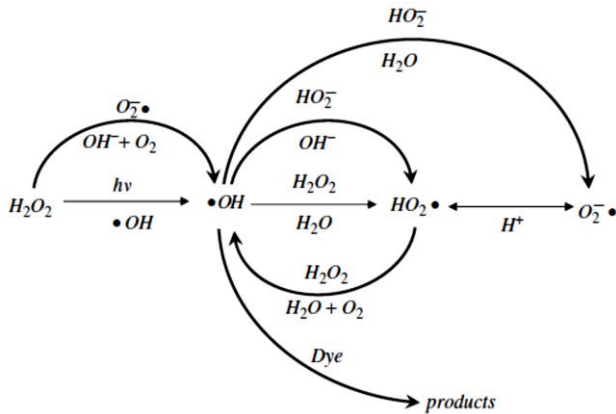
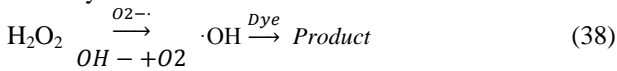


Figure 1. The reaction network for azo dye degradation using UV/ H_2O_2 as proposed by Chang et al. (2010).

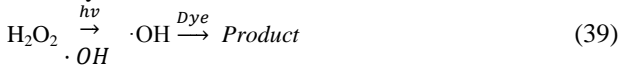
The following dye decomposition rate was derived for three main pathways, assuming that the pseudo-steady state can be used for hydroxyl and hydroperoxyl free radicals:

$$-r_{dye} = r_1 + r_2 + r_3 \quad (37)$$

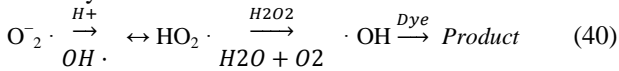
Pathway 1:



Pathway 2:



Pathway 3:



$$-r_{dye0} = \frac{k_{o1}k_aI_0 \exp(-k_H[H_2O_2] - k_D[Dye])[H_2O_2][Dye](1+k_b[H_2O_2])}{[H_2O_2] + k_a[Dye] + k_b[H_2O_2]} \quad (41)$$

In comparison to other studies that use a pseudo-first-order kinetic model, that study specifies the individual reaction rate coefficient between the azo dye and the hydroxyl radical (Chang et al., 2010). Studies reviewed in this section indicate that the azo dye degradation process can be successfully approximated with a first-order decay model. In the studies reviewed, different explanations for the decay rate constant are developed to come up with a better explanation for the dye degradation process. In the following section, a mathematical model for a reactor that is used for the photocatalytic dye degradation process is explained.

Reactor Design Model

Shen and Wang (2002) developed a model for photoreactor design for azo dye decomposition with

UV/ H_2O_2 . That model can be used to predict the dye degradation within different type (geometry) of photoreactors under various operating conditions. The three-dimensional transport equation in the UV/ H_2O_2 system is shown as follows in cylindrical coordinates with a constant density (ρ) and diffusivity (D_{AB}):

$$\frac{\partial C_{dye}}{\partial t} + \left(v_r \frac{\partial C_{dye}}{\partial r} + v_\theta \frac{1}{r} \frac{\partial C_{dye}}{\partial \theta} + v_z \frac{\partial C_{dye}}{\partial z} \right) = D_{AB} \cdot \left(\frac{1}{r} \frac{\partial}{\partial r} \left(r \frac{\partial C_{dye}}{\partial r} \right) + \frac{1}{r^2} \frac{\partial^2 C_{dye}}{\partial \theta^2} + \frac{\partial^2 C_{dye}}{\partial z^2} \right) - R_{dye} \quad (42)$$

where C_{dye} is dye concentration (mM), D_{AB} the binary diffusivity of dye in water (l^2s^{-1}), and R_{dye} is the rate of dye photooxidation ($M s^{-1} L^{-1}$).

Equation 42 was simplified under the following assumptions (Shen and Wang 2002):

1. Accumulation variability is not considered as the system is assumed to be steady-state.
2. Because an ideal plug flow is assumed, then $v_r = 0$, $v_\theta = 0$, and $v_z = V/\tau$ constant, where V is the annular reactor space volume, τ is the dye-containing solution retention time within the photoreactor, v_z is the solution velocity in the z -direction.
3. The diffusion terms are considered to be insignificant compared to the convection terms.
4. The reaction intermediates produced during the photooxidation will not interfere with the parent dye's photodegradation.

Therefore, the Equation 42 is simplified to:

$$0 + v_z \frac{\partial C_{dye}}{\partial z} = -R_{dye} \quad (43)$$

The design equation can be written as follows for a plug-flow reactor:

$$\tau = \frac{V}{v_z} = - \int_{C_{dye,in}}^{C_{dye,eff}} \frac{dC_{dye}}{R_{dye}} \quad (44)$$

The photooxidation rate (R_{dye}) is a function of the UV light intensity (I) and the H_2O_2 dosage:

$$-R_{dye} = k \cdot I^m \cdot [H_2O_2]^n \cdot C_{dye}^p \quad (45)$$

where k is the rate constant, I is the UV light intensity (Wm^{-2}), $[H_2O_2]$ the H_2O_2 dosage ($mM l^{-1}$) and m , n , p are the orders with respect to I , H_2O_2 dosage and C_{dye} and can be experimentally determined.

UV irradiance, I , at any point inside the reactor is related to the surface flux, I_0 , in the infinite line source model. Using a one-dimensional Lambert's law the functional relationship between UV irradiance and surface flux can be given as follows:

$$I = I_0 \cdot \frac{r_i}{r} e^{-E(r-r_i)} \quad (46)$$

where E is the monochromatic absorbance of water. E value is determined to be 0.624 cm^{-1} in that work.

Equation 46 is modified with the dye and H_2O_2 in the aqueous stream as:

$$I = I_0 \cdot \frac{r_i}{r} e^{-(E + \epsilon_{dye} C_{dye} + \epsilon_{H_2O_2}) \cdot (r-r_i)} \quad (47)$$

where the ε_{dye} and $\varepsilon_{\text{H}_2\text{O}_2}$ are the molar absorption coefficients of the dye and H_2O_2 at 254 nm UV light and were determined as 12.716 and $19\text{M}^{-1}\text{cm}^{-1}$, respectively.

The terms $\varepsilon_{\text{dye}}C_{\text{dye}}$ and $\varepsilon_{\text{H}_2\text{O}_2}C_{\text{H}_2\text{O}_2}$ shown in Equation 47 represent the radiation decrease caused by dye and H_2O_2 absorption. Combining Equations 43–45 and 47 yields:

$$\frac{\partial C_{\text{dye}}}{\partial \tau} = \left(v_z \frac{\partial C_{\text{dye}}}{\partial z} \right) = -R_{\text{dye}} = k \left[I_0 \cdot \frac{r_i}{r} e^{-(E + \varepsilon_{\text{dye}} C_{\text{dye}} + \varepsilon_{\text{H}_2\text{O}_2} C_{\text{H}_2\text{O}_2}) \cdot (r - r_i)} \right]^m \cdot [\text{H}_2\text{O}_2]^n \cdot C_{\text{dye}}^p \quad (48)$$

Integration of Eq. (48) for the whole length of the reactor gives:

$$\int_{C_{\text{dye},\text{in}}}^{C_{\text{dye},\text{eff}}} \frac{dC_{\text{dye}}}{C_{\text{dye}}^p e^{-\varepsilon_{\text{dye}} C_{\text{dye}} m (r - r_i)}} = k \left[I_0 \cdot \frac{r_i}{r} e^{-(E + \varepsilon_{\text{H}_2\text{O}_2} C_{\text{H}_2\text{O}_2) \cdot (r - r_i)} \right]^m \cdot [\text{H}_2\text{O}_2]^n \cdot L / v_z \quad (49)$$

where $C_{\text{dye},\text{in}}$ and $C_{\text{dye},\text{eff}}$ are the influent and effluent concentrations of dye (mM). The integrated value of $C_{\text{dye},\text{eff}}(r)$ depends on radial distance, r , since UV light attenuates as radial distance increases.

The average residual concentration of dye, $C_{\text{dye},\text{eff},\text{ave}}$, in the photoreactor effluent is calculated with the following equation.

$$C_{\text{dye},\text{eff},\text{ave}} = \frac{\int_0^{2\pi} \int_{r_i}^{r_0} C_{\text{dye},\text{eff}}(r) dr d\theta}{\int_0^{2\pi} \int_{r_i}^{r_0} (r) dr d\theta} = C_{\text{dye},\text{eff},\text{ave}} \left(I_0, C_{\text{dye}}, v_z, \frac{r_i}{r_0} L, k, m, n, p \right) \quad (50)$$

Equation 50 is the design equation for dye photodegradation in the annular reactor and in that equation I_0 , C_{dye} , v_z are operating parameters, r_i/r_0 , L are

geometrical parameters, and k , m , n , p are reaction parameters (Shen and Wang, 2002). The study by Shen and Wang (2002) is a good example of combining both empirical and mathematical models. Analysis of mathematical models starts with the unrealistic steady-state assumption in most cases and the solution of mathematical models including partial differential equations are difficult and only numerically feasible.

In the next section, studies utilized RSM to estimate azo dye degradation by UV/ H_2O_2 process are reviewed.

Response Surface Models

The structures of a first and second-order response surface models are given in Equations 51 and 52 below:

$$y = \beta_0 + \beta_1 x_1 + \beta_2 x_2 + \dots + \beta_k x_k + \varepsilon \quad (51)$$

$$y = \beta_0 + \sum_{i=1}^k \beta_i x_i + \sum_{i=1}^k \beta_{ii} x_i^2 + \sum_{i < j} \beta_{ij} x_i x_j + \varepsilon \quad (52)$$

where stands for predicted response, β_0 is intercept, β_i first order term for factor i , β_{ii} is the second order term for factor i and, represents the interaction effect. x_i and x_j are factor values in the experiment (e.g., pH, temperature).

The interaction term is a measure of how much the dependent variable, with respect to one factor, changes as the other factor increases or decreases (Palasota et al., 1992).

Response surface methodology has recently been successfully applied to different processes for achieving their optimization (Savun-Hekimoğlu and Ince 2019). However, the application of RSM to a UV/ H_2O_2 process has not been widely reported. Table 2 summarizes the relevant literature on RSM studies for azo dye degradation by UV/ H_2O_2 .

Table 2. RSM studies in the literature to evaluate degradation of azo dyes by UV/ H_2O_2 process.

Response	Variables	Optimum conditions	References
Decolorization Dearomatization	initial dye concentration (c_{d0}), initial H_2O_2 concentration (c_{h0}), pH	Decolorization: c_{d0} : 250 mg L ⁻¹ c_{h0} : 30 mM pH: 7.30 Dearomatization: c_{d0} : 250 mg L ⁻¹ c_{h0} : 30 mM pH: 6.87	Zuorro et al., 2013
Decolorization	Dye (mM), H_2O_2 (M), pH	0.005mM dye, 0.042M H_2O_2 , 6.6 pH	Rauf et al., 2008
Decolorization	Initial concentration of Dye-1 (mg/L), Dye-2 (mg/L), Dye-3 (mg/L), initial H_2O_2 concentration (mg/L), reaction time (min)	4 mg/L dye (three dyes' initial concentrations), 48 mg/L H_2O_2 and 30 min	Khataee et al., 2012
Decolorization	initial dye concentration, H_2O_2 concentration, reaction time and distance from UV lamp	20 mg/L initial dye concentration, 1 g/L H_2O_2 concentration, 14 min reaction time, and 16 cm distance from the UV lamp.	Khataee and Habibi, 2010
Mineralization (TOC% decay)	Temperature (°C), H_2O_2 (mM), Time (min)	217 °C, 111 mM, and 60 min	Kayan and Gözmen, 2012

In the studies summarized in Table 2, the RSM approach provided useful indications on the individual and interactive effects of the process variables on azo dye degradation via UV/H₂O₂. Using RSM valid regression models were developed which provide reliable predictions of the performance of UV/H₂O₂ reaction systems, using experimental data generated under the specified operating conditions (e.g. reactor, reagent and ambient conditions).

However, in most of those studies, optimum operating conditions were determined visually from 3-D response surface plots. Also, a sensitivity analysis was not conducted in those studies. In order to fill this gap in the literature, at the end of the following section, a RSM application is conducted for dye decolorization with the UV/H₂O₂ process.

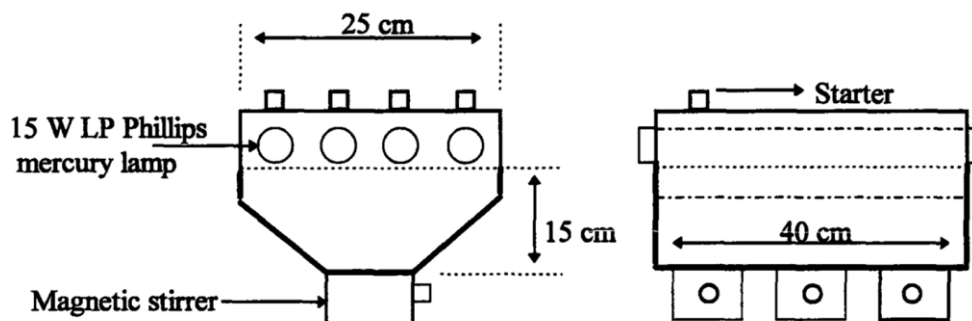


Figure 2. The schematic diagram of the photoreactor (Ince and Apikyan, 2000).

First-order decay is assumed for dye degradation in our analysis. Using Equation 2 given in the first section, reaction constant k is calculated for each experiment set. These values are given in Table 3.

Table 3. Pseudo-first-order rate constants of the azo dyes photodegradation.

k (min ⁻¹)	pH	[C _{dye}] ₀
0.224369	3	36
0.219239	4.5	36
0.192919	7	36
0.240024	9.5	36
0.102029	11	36
0.685631	7	9
0.495564	7	18
0.252367	7	36
0.104741	7	72

Before developing our regression model by using RSM, first, the equations given by Behnajady et al. (2006) are applied to the observed data set in order to verify the coefficient values. Using Equations 29-30 reaction rate constant (k) is estimated as $k = 0.884494 \text{ min}^{-1}$. However, the observed data shows that the value of the reaction rate constant is 0.252367. Obviously, the observed and predicted k values do not match with each other.

Black-box modeling applications with the experimental data

In this section of this paper, some of the models reviewed in the previous sections are applied to a set of data achieved through laboratory experiments. Reactive Black 5 (RB5) is selected as the model compound and the experimental set-up used in the study is presented in Figure 1. Four UV lamps (Philips 15-W low-pressure mercury), emitting monochromatic light at 253.7 nm were used. The distance between the surface of the solution and the UV lamps was 6 cm. In order to determine the optimum pH in this system, 36 ppm dye was exposed to UV irradiation in the presence of 400 mg/L H₂O₂ at pH 3, 4.5, 7, 9.5, and 11. The effect of initial dye concentration on color removal was determined by exposing 18, 36, 45 and 72 ppm dye to UV/H₂O₂ oxidation at pH 7. The color analyses were carried out spectrophotometrically.

In addition to that, the equation given by Malik and Sanyal (2004) is applied to the experimentally achieved dataset as shown in Eq. (53).

$$k = 8.11[\text{H}_2\text{O}_2]_0^{0.63} = 8.11(0.012)^{0.63} = 0.5 \text{ min}^{-1} \quad (53)$$

It is obvious from the Equation (53) and Table 3 that the calculated result is higher than the observed k values. The difference between calculated and observed k values can be attributed to the use of different azo dyes having different chemical structures or can be attributed to different reagent doses and operating conditions. These findings indicate the need for recalculation of reaction orders and coefficient values. More accurate estimations of the coefficients and a detailed study on the comparison of the kinetic models with a wider range of experimental data are left to future research.

Although empirical models based on experimental findings exist in the literature for the UV/H₂O₂ process, the effect of pH is still unclear and needs to be explained in detail. In this part of this study, it is aimed to fill this gap by using RSM for the data set given in Table 3. The reactions were run in the photoreactor at various preset values of pH and initial dye concentrations, which were chosen to be the independent variables of the model. The dependent (response) variable reflecting the efficiency of the system was selected as the first-order reaction rate constant.

In this study, the size of the data set is nine whereas the number of experimental parameters is two (pH and initial dye concentration). First-order terms are included in the model as a first step. Coefficients of this model are calculated in SPSS which is a common statistics software. Summary statistics for this model are given in Table 4. Since the R-square value is found to be 0.565, second-order terms are included in the model which is given in Equation 54.

Table 4. Summary statistics for first regression model.

Model Summary				
Model	R	R Square	Adjusted R Square	Std. Error of the Estimate
1	,821 ^a	,673	,565	,12567

a. Predictors: (Constant), C_0, pH

$$y = b_0 + b_1x_1 + b_2x_2 + b_3x_1^2 + b_4x_2^2 \quad (54)$$

Montgomery (2001) states that second-order response surface models should include interaction terms as well as squares of each independent variable as shown in Equation 54. However, adding interaction terms to Equation 54 reveal that our data set is not suitable for

this calculation. Therefore, RSM with a model including the interaction term is left to future research due to the requirement of additional experiments. In this paper model including only first and second-order terms without interactions is used for RSM. SPSS output including summary statistics for this model is given in Table 5. The R-square value of the model indicates that this polynomial function provides a good fit to the data set.

Table 5. Summary statistics for second regression model.

Model Summary ^b				
Model	R	R Square	Adjusted R Square	Std. Error of the Estimate
1	,989 ^a	,978	,956	,03997

a. Predictors: (Constant), C0_2, pH, C_0, ph2

b. Dependent Variable: k

SPSS output for coefficients of the response surface model is given in Table 6. By substituting these coefficient values in Equation 54 response surface model given in Equation 55 is derived.

Table 6. Regression coefficients of the response surface model.

Coefficients ^a						
Model		Unstandardized Coefficients		Standardized Coefficients	t	Sig.
		B	Std. Error	Beta		
1	(Constant)	,777	,153		5,066	,007
	pH	,045	,035	,559	1,305	,262
	C_0	-,026	,003	-,332	-,781	,001
	ph2	-,004	,002	-,695	-,614	,182
	C0_2	,000	,000	1,597	5,372	,006

a. Dependent Variable: k

$$y = 0.77 + 0.0451242x_1 - 0.025968x_2 - 0.003952x_1^2 + 0.0002073x_2^2 \quad (55)$$

3-D plot for this response surface, which is given in Figure 3, is obtained from MATLAB. The contour plot for this surface is also given in Figure 4. Obviously, k increases steeply for low values of initial dye concentration. As one can see, there is an inverse relationship between the initial dye concentration and the reaction rate constant. Galindo and Kalt (1999) explained that an inverse relationship with an increase in the dye concentration induces a rise in the internal optical density. As the dyes react the impermeability of the solution to UV radiation increases. Therefore, the dependency of the H₂O₂-photolysis-rate to the incident intensity was reported. As a consequence, with the increasing concentrations of azo dye, the hydroxyl radicals generation of the system decreases (Muruganandham et al., 2014).

One common usage of RSM is the optimization of experiment output with respect to its parameters (Montgomery, 2001).

Since the reaction rate constant is the primary determinant for the duration of dye degradation, the maximization of k value is a significant task for dye removal processes. In this study, the response surface model given above is used for calculating the optimum pH value which maximizes the reaction rate constant. Optimum pH value that maximizes reaction constant is found by taking the first derivative and having it equal to zero (Montgomery, 2001). This calculation shows that the optimum pH value that gives the maximum reaction constant is 5.74. Another important issue in chemical reactions is finding the most effective parameter that can be used to control the process. For this purpose, different sensitivity analysis methods are proposed in the literature (Hamby, 1994). Among these proposed methods standardized regression coefficient (SRC) is selected for analysis of the sensitivity of rate constant to experiment parameters which are initial dye concentration (C₀) and pH. In this method, a regression equation, which minimizes the sum of squares of residual terms, is calculated by using the ordinary least squares algorithm. The use of the regression technique allows the sensitivity ranking to be determined based on the relative magnitude of the regression coefficients (Hamby, 1994). The coefficients indicate the amount of influence the parameter has on the model as a whole.

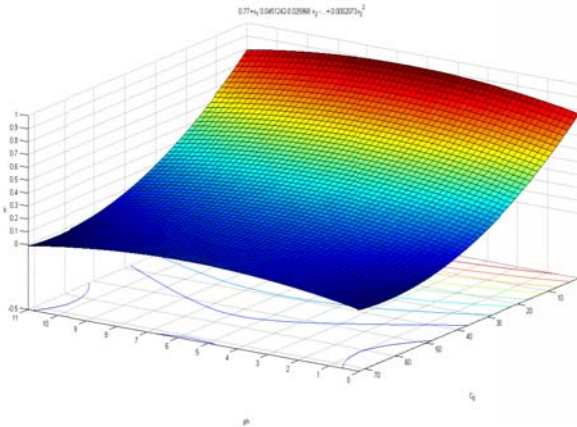


Figure 3. 3-D Plot of Response Surface for Reaction Constant *k*.

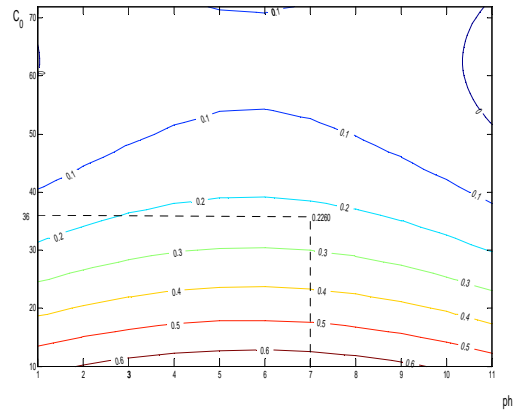


Figure 4. Contour Plot of Response Surface for Reaction Constant *k*.

Table 6. Regression coefficients of the first-order regression model.

Coefficients^a

Model		Unstandardized Coefficients		Standardized Coefficients	t	Sig.
		B	Std. Error	Beta		
1	(Constant)	,665	,166		4,015	,007
	pH	-,010	,019	-,122	-,522	,621
	C_0	-,009	,003	-,812	-3,478	,013

a. Dependent Variable: *k*

In a standard regression model, given in Equation 55, the dependent variable (*y*) is tried to be explained by *k* independent variables (*x_i*). Standardized regression coefficients, given in Equation 56, give the importance of independent variables for the dependent one in a regression equation (Saltelli et al., 2000). In other words, the output is more sensitive to the parameters that have larger-magnitude standardized regression coefficients in the regression equation.

$$y = \beta_0 + \beta_1 x_1 + \beta_2 x_2 + \dots + \beta_k x_k + \varepsilon \quad (55)$$

$$b_j = \frac{\beta_j \sigma_{x_j}}{\sigma_y} \quad (56)$$

In this study, experiment data, given in Table 3, is entered into SPSS and a multi-variable linear regression model is calculated of which SPSS outputs are given in Table 6.

SRC of pH is found to be -0.122 whereas SRC of initial dye is calculated -0.812. According to these results, initial dye concentration is a more significant parameter for reaction constant *k*. This conclusion can also be seen from the response surface given in Figure 3. Note that in the kinetic model of El-Dein et al. (2001) the impact of initial dye concentration was neglected. However, our findings show that the impact of initial dye concentration is not negligible. In agreement with our results, Shu et al. (1994) and Galindo and Kalt (1999) found that the decolorization rate depends on the molar ratio of H₂O₂ to the dye concentration, pH and chemical structure of the dye.

Conclusions

Azo dye decolorization process is studied by various researchers with empirical, semi-empirical and mathematical models. In general, experiments indicate that the azo dye degradation process can be well approximated by first-order decay model. Different studies that provide various explanations to the decay rate constant can be found in the literature. Some of the researchers consider the combined effect of UV and H₂O₂ in the decay rate constant whereas others also consider the respective effects of UV and H₂O₂ separately in an additive manner. Although numerous researchers have come up with various expressions for predicting decay rate constant *k*, obviously there is no well-established model that predicts *k* value under different operating conditions. Moreover, empirical studies with response surface models, which are limited in number, are conducted to suggest the optimum operating conditions. In this study, a review of models in the literature is followed by the application of some models to a data set achieved through laboratory experiments. Calculations indicate that estimated values are slightly different than the observed decay rate constant values. This finding is attributed to different chemical structures of different azo dyes and different operating conditions.

In the next step of this study, Response Surface Methodology (RSM) is utilized in order to approximate the functional relationship between *k*, pH and initial dye

concentration. A second-order polynomial function without interaction terms is assumed for the response surface model. Coefficients of this polynomial are calculated using the least square method in SPSS. The high R-square value of this response surface model indicates that the second-order polynomial function is a proper approximation. In this paper, RSM is used to maximize the decay rate constant which plays a significant role in the decolorization process. Using this response surface model optimum pH, maximizing the decay rate constant is calculated. The calculated optimum pH is in agreement with the literature.

In the last step sensitivity analysis using standardized regression coefficients (SRCs) is conducted for data set. A linear first-order regression model is assumed and SRCs are calculated using SPSS. The result of sensitivity analysis indicates that initial dye concentration is a more significant parameter for decay rate constant than pH value. This finding is in complete agreement with the literature reviewed above.

For further research, optimization with a response surface model including interaction terms is suggested. For that purpose, additional experiments need to be conducted with various parameter values. Also, toxicity is not analyzed or considered in any of the reviewed studies. Hence, an RSM study considering the relationship between the toxicity of the generated byproducts and the operating parameters such as pH, H₂O₂ concentration and most importantly irradiation time seems to be necessary.

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