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## Articles and Statements

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### Can the First Organic Forms Of Life Originate in Hot Mineral Water with HDO?

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#### Abstract

The isotopic composition, the temperature and the pH value of water were analyzed in experiments with modelling the primary hydrosphere and hydrological conditions for origin of first organic forms in hot mineral water with HDO. For this aim we performed experiments with hot mineral water and seawater from Bulgaria and water with varying content of deuterium using IR- and DNES-spectroscopy. As model systems were used cactus juice of *Echinopsis pachanoi* and Mediterranean jellyfish *Cotylorhiza tuberculata*. The reactions of condensation–dehydration occurring in alkaline aqueous solutions at  $t = 65\text{--}95\text{ }^{\circ}\text{C}$  and  $\text{pH} = 9\text{--}10$ , resulting in synthesis from unorganic molecules the larger organic molecules as polymers and short polipeptides, were discussed, as well as the possible mechanisms of the deuterium accumulation in form of HDO in hot water. It was shown that hot alkaline mineral water with temperature from  $+65\text{ }^{\circ}\text{C}$  to  $+95\text{ }^{\circ}\text{C}$  and the pH value from 9 to 11 is more suitable for the origination of life and living matter than other analyzed water samples. In hot mineral waters the local maximums in IR-spectra are more manifested compared to the local maximums obtained in IR-spectra of the same water at a lower temperature. The difference in the local maximums from  $+20\text{ }^{\circ}\text{C}$  to  $+95\text{ }^{\circ}\text{C}$  at each  $+5\text{ }^{\circ}\text{C}$  according to the Student  $t$ -criterion makes up  $p < 0,05$ .

**Keywords:** deuterium, hydrosphere, evolution, origin of life, IR spectroscopy, DNES.

#### Introduction

The previous biological experiments with  $\text{D}_2\text{O}$  and structural-conformational studies with deuterated macromolecules, performed by us, enable to modeling conditions under which the first living forms of life might be evolved [1–3]. The content of deuterium in hot mineral water may be increased due to the physical-chemical processes of the deuterium accumulation. It can be presumed that primary water might contain more deuterium at early stages of evolution of first

living structures, and deuterium was distributed non-uniformly in the hydrosphere and atmosphere [4]. The primary reductive atmosphere of the Earth consisted basically of gas mixture CO, H<sub>2</sub>, N<sub>2</sub>, NH<sub>3</sub>, CH<sub>4</sub>, lacked O<sub>2</sub>–O<sub>3</sub> layer protecting the Earth surface from rigid short-wave solar radiation carrying huge energy capable to cause radiolysis and photolysis of water. The processes accompanying accumulation of deuterium in the hydrosphere are solar radiation, volcanic geothermal processes and electric discharges in the atmosphere. These natural processes could lead to the enrichment of the hydrosphere by deuterium in the form of HDO which evaporates more slowly than H<sub>2</sub>O, and condenses faster. If this is true, this is a significant fact regarding thermal stability of deuterated macromolecules in the preservation of life under thermal conditions, because chemical bonds with participation of deuterium are stronger than those ones formed of hydrogen.

The natural prevalence of deuterium makes up approximately 0,015–0,020 atom%, and depends strongly on the uniformity of substance and the total amount of matter formed in the course of early Galaxy evolution [5]. The average ratio of D/<sup>1</sup>H in nature makes up approximately 1:5700. In natural waters, the deuterium is distributed irregularly: from 0,02–0,03 mol.% for river water and sea water, to 0,015 mol.% for water of Antarctic ice – the most purified from deuterium natural water containing in 1,5 times less deuterium than that of seawater. According to the international SMOW standard isotopic shifts for D and <sup>18</sup>O in sea water: D/<sup>1</sup>H = (155,76±0,05)·10<sup>-6</sup> (155,76 ppm) and <sup>18</sup>O/<sup>16</sup>O = (2005,20±0,45)·10<sup>-6</sup> (2005 ppm). For the SLAP standard isotopic shifts for D and <sup>18</sup>O in seawater make up D/<sup>1</sup>H = 89·10<sup>-6</sup> (89 ppm) and for a pair of <sup>18</sup>O/<sup>16</sup>O = 1894·10<sup>-6</sup> (1894 ppm). In surface waters, the ratio D/<sup>1</sup>H = ~(1,32–1,51)·10<sup>-4</sup>, while in the coastal seawater – ~(1,55–1,56)·10<sup>-4</sup>. The natural waters of CIS countries are characterized by negative deviations from SMOW standard to (1,0–1,5)·10<sup>-5</sup>, in some places up to (6,0–6,7)·10<sup>-5</sup>, but however there are also observed positive deviations at 2,0·10<sup>-5</sup>.

The constant sources of deuterium are explosions of nova stars and thermonuclear processes frequently occurring inside the stars. Probably, it could explain a known fact, why the amount of deuterium is slightly increased during the global changes of climate in warming conditions. The gravitational field of the Earth is insufficiently strong for the retaining of lighter hydrogen, and our planet is gradually losing hydrogen as a result of its dissociation into interplanetary space. Hydrogen evaporates faster than heavy deuterium, which can be collected by the hydrosphere. Therefore, as a result of this natural process of fractionation of H/D isotopes throughout the process of Earth evolution there should be an accumulation of deuterium in the hydrosphere and surface waters, while in the atmosphere and in water vapour deuterium content tends to be low. Thus, on the planet there occurs a natural process of separation of H and D isotopes, playing an essential role in the maintenance of life on the planet.

The second point regards the influence of temperature on the biochemical processes in living matter. Our recent studies have shown that the most favorable for the origin of life and living matter seem to be hot alkaline mineral waters interacting with CaCO<sub>3</sub> [6, 7]. According to the law for conservation of energy the process of self-organization of primary organic forms in water solutions may be supported by the thermal energy of magma, volcanic activity and solar radiation. According to J. Szostak, the accumulation of organic compounds in small isolated lakes is more possible compared to the ocean [8]. It is most likely that life originated near a hydrothermal vent: an underwater spout of hot water. The geothermal activity gives more opportunities for the origination of life. In 2009 A. Mulkidjanian and M. Galperin demonstrate that the cell cytoplasm contains potassium, zinc, manganese and phosphate ions, which are not particularly widespread in the sea aquatorium [9]. J. Trevors and G. Pollack proposed in 2005 that the first cells on the Earth assembled in a hydrogel environment [10]. Gel environments are capable of retaining water, oily hydrocarbons, solutes, and gas bubbles, and are capable of carrying out many functions, even in the absence of a membrane. Hydrocarbons are an organic compounds consisting entirely of hydrogen and carbon. The data presented in this paper show that the origination of living matter most probably occurred in hot mineral water. This may occurred in ponds and hydrothermal vents in seawater or hot mineral water. An indisputable proof of this is the presence of stromatolites fossils. They lived in warm and hot water in zones of volcanic activity, which could be heated by magma and seem to be more stable than other first sea organisms [11].

Therefore, the purpose of the research was studying the hydrological conditions of primary hydrosphere (temperature, pH, isotopic composition) for possible processes for origin of first organic forms in hot mineral water with HDO. Various samples of water from Bulgaria, as well water with varying deuterium content were studied within the frames of the research.

## **Material and Methods**

### **Chemicals**

For preparation of water with varying content of deuterium (HDO) was used D<sub>2</sub>O (99,9 atom.%) from the Russian Research Centre “Isotope” (St. Petersburg, Russian Federation). D<sub>2</sub>O was preliminary distilled over KMnO<sub>4</sub> with the subsequent control of isotope enrichment by <sup>1</sup>H-NMR-spectroscopy on a Bruker WM-250 device (“Bruker”, Germany) (working frequency: 70 MHz, internal standard: Me<sub>4</sub>Si).

### **Biological Objects**

The objects of the study were used the cactus juice of *Echinopsis pachanoi* and the Mediterranean jellyfish *Cotylorhiza tuberculata* (Chalkida, Greece, Aegean Sea).

### **Water Samples**

The samples of water were taken from various water springs of Bulgaria:

- 1 – Mineral water (Rupite, Bulgaria);
- 2 – Seawater (Varna resort, Bulgaria);
- 3 – Mountain water (Teteven, Bulgaria);
- 5 – Deionized water (the control).
- 6 – Water with varying deuterium content (HDO).

### **IR-Spectroscopy**

IR-spectra of water samples were registered on Bruker Vertex (“Bruker”, Germany) Fourier-IR spectrometer (spectral range: average IR – 370–7800 cm<sup>-1</sup>; visible – 2500–8000 cm<sup>-1</sup>; permission – 0.5 cm<sup>-1</sup>; accuracy of wave number – 0.1 cm<sup>-1</sup> on 2000 cm<sup>-1</sup>) and on Thermo Nicolet Avatar 360 Fourier-transform IR (M. Chakarova).

### **DNES-Spectral Analysis**

The research was made with the method of differential non-equilibrium spectrum (DNES). The device measures the angle of evaporation of water drops from 72 ° to 0 °. As the main estimation criterion was used the average energy ( $\Delta E_{H...O}$ ) of hydrogen O...H-bonds between H<sub>2</sub>O molecules in water’s samples. The spectra of water were measured in the range of energy of hydrogen bonds 0,08–0,1387 eV with using a specially designed computer program.

### **High-Frequency Coronal Electric Discharge Experiments**

A device for high-frequency coronal electric discharge was used in this study, constructed by I. Ignatov and Ch. Stoyanov. The frequency of the applied saw-tooth electric voltage was 15 kHz, and the electric voltage – 15 kV. The electric discharge was obtained using a transparent firm polymer electrode on which a liquid sample of water (2–3 mm) was placed. The spectral range of the photons released upon electric discharge was from  $\lambda = 400$  to  $\lambda = 490$  nm and from  $\lambda = 560$  to  $\lambda = 700$  nm.

## **Results and Discussion**

### **Studying Various Water Samples on the Feasibility for Origin of Life**

We have carried out the research of various samples of mineral water obtained from mineral springs and seawater from Bulgaria (Fig. 1, curves 1–5). For this aim we employed IR-spectrometry and the DNES method relative to the control – deionized water.

For calculation of the function  $f(E)$  represented the energy spectrum of water, the experimental dependence between the wetting angle ( $\theta$ ) and the energy of hydrogen bonds ( $E$ ) is established:

$$f(E) = \frac{14,33f(\theta)}{[1-(1+bE)^2]^2}, \quad (1)$$

where  $b = 14,33 \text{ eV}^{-1}$

The relation between the wetting angle ( $\theta$ ) and the energy ( $E$ ) of the hydrogen bonds between  $\text{H}_2\text{O}$  molecules is calculated by the formula:

$$\theta = \arcsin(-1 - 14,33E) \quad (2)$$

Cactus juice was also investigated by the DNES method (Fig. 1, curve 1). The cactus was selected as a model system because this plant contains approximately 90 % of water. The closest to the spectrum of the cactus juice was the spectrum of mineral water contacting with  $\text{Ca}^{2+}$  and  $\text{HCO}_3^-$  ions (Fig. 1, curve 2). DNES-spectra of cactus juice and mineral water have magnitudes of local maximums ( $E$ , eV) at  $-0,1112$ ;  $-0,1187$ ;  $-0,1262$ ;  $-0,1287$  and  $-0,1387$  eV. The similar local maximums in the DNES-spectrum between the cactus juice and seawater were detected at  $-0,1362$  eV. The DNES-spectrum of the control sample of deionized water (Fig. 1, curve 5) was substantially different from DNES-spectra of seawater and mineral water.

Another important parameter was measured by the DNES method – the average energy ( $\Delta E_{\text{H...O}}$ ) of hydrogen  $\text{H...O}$ -bonds among the individual molecules  $\text{H}_2\text{O}$ , which makes up  $-0,1067 \pm 0,0011$  eV. When the water temperature is changed, the average energy of hydrogen  $\text{H...O}$ -bonds alternates. This testified about the restructuring of average energies of hydrogen  $\text{H...O}$ -bonds among individual  $\text{H}_2\text{O}$  molecules with a statistically reliable increase of local maximums in DNES-spectra.

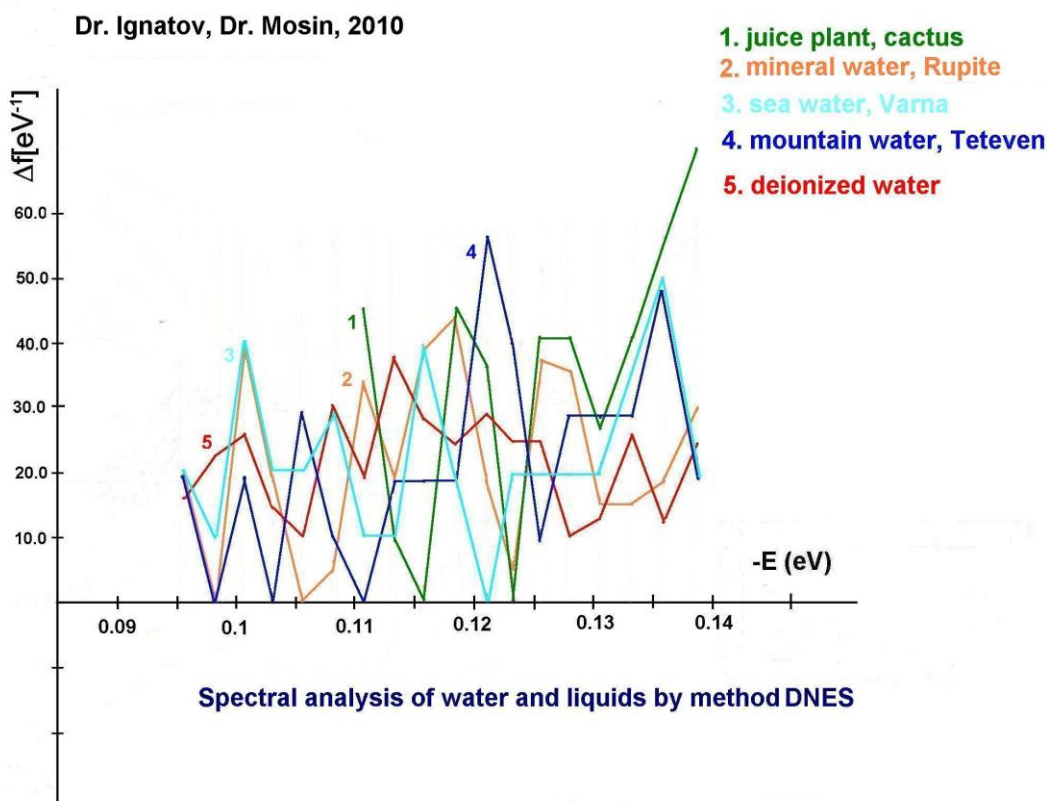


Figure 1. DNES-spectra of water samples of various origin: 1 – the cactus juice; 2 – mineral water from Rupite village (Bulgaria); 3 – seawater (Varna, Bulgaria); 4 – mountain water (Teteven, Bulgaria); 5 – deionized water (the control)

As it was shown from these experimental data, the closest to the IR-spectrum of the cactus juice was mineral water from Rupite Village (Bulgaria), which the DNES and IR spectrum are shown in Fig. 2 and Fig. 3 (Thermo Nicolet Avatar 360 Fourier-transform IR). IR-spectra of cactus juice and mineral water containing  $\text{HCO}_3^-$  (1320–1488 mg/l),  $\text{Ca}^{2+}$  (29–36 mg/l), pH (6,85–7,19), have local maximums at  $\lambda = 8,95; 9,67; 9,81; 10,47$  and  $11,12 \mu\text{m}$  (Fourier-IR spectrometer Brucker Vertex). The common local maximums in the IR-spectrum between the cactus juice and seawater are detected at  $\lambda = 9,10 \mu\text{m}$ . The local maximums obtained with the IR method at  $\lambda = 9,81 \mu\text{m}$  ( $k = 1019 \text{ cm}^{-1}$ ) and  $\lambda = 8,95 \mu\text{m}$  ( $k = 1117 \text{ cm}^{-1}$ ) (Thermo Nicolet Avatar 360 Fourier-transform IR) are located on the spectral curve of the local maximum at  $\lambda = 9,7 \mu\text{m}$  ( $k = 1031 \text{ cm}^{-1}$ ) (Fig. 3). With the DNES method, the results of which are shown in Table 1, were obtained the following results – 8,95; 9,10; 9,64; 9,83; 10,45 and 11,15  $\mu\text{m}$  (wave length  $\lambda$ ,  $\mu\text{m}$ ), or – 897; 957; 1017; 1037; 1099 and  $1117 \text{ cm}^{-1}$  (wave numbers,  $k$ ,  $\text{cm}^{-1}$ ).

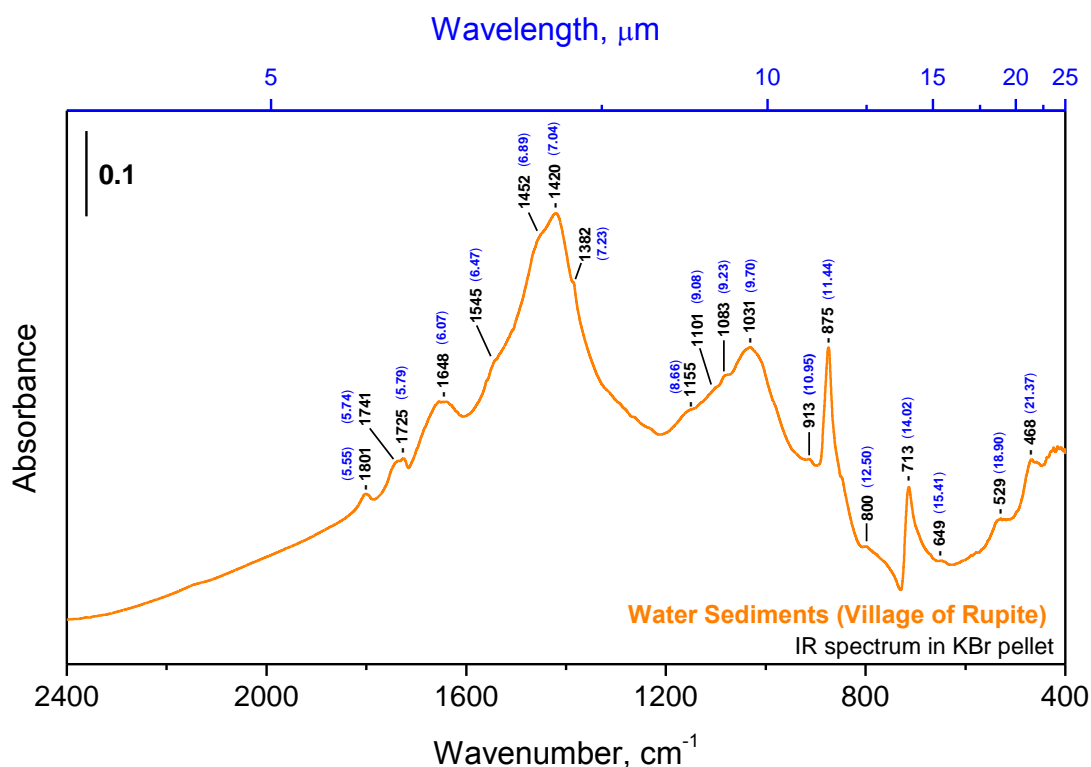


Figure 2. IR-spectrum of water sediments obtained from Rupite Village (Bulgaria)

Table 1: Characteristics of spectra of water of various origin obtained by the DNES-method\*

-E, eV			$\lambda$ , $\mu\text{m}$	$k$ , $\text{cm}^{-1}$
Cactus juice	Mineral water from Rupite Village (Bulgaria)	Seawater		
0,1112	0,1112	–	11,15	897
0,1187	0,1187	–	10,45	957
0,1262	0,1262	–	9,83	1017
0,1287	0,1287	–	9,64	1037
0,1362	–	0,1362	9,10	1099
0,1387	0,1387	–	8,95	1117

Notes:

\*The function of the distribution of energies  $\Delta f$  was measured in reciprocal electron volts ( $\text{eV}^{-1}$ ). It is shown at which values of the spectrum  $-E$  (eV) were observed the biggest local maximums of this function;  $\lambda$  – wave length;  $k$  – wave number.

The results with the Mediterranean jellyfish *Cotylorhiza tuberculata* indicated that the jellyfish has local maximums in IR-spectra at  $\lambda = 8,98$  and  $\lambda = 10,18 \mu\text{m}$  (Fig. 3). Before measurements the jellyfish was kept in seawater for several days. On comparison the seawater has a local maximum at  $\lambda = 8,93 \mu\text{m}$  in IR-spectra. These results were obtained with Thermo Nicolet Avatar 360 Fourier-transform IR. With the DNES method the local maximums in spectra for jellyfish are detected at  $\lambda = 8,95$  and  $10,21 \mu\text{m}$ , and for seawater – at  $\lambda = 9,10 \mu\text{m}$ . A differential spectrum was recorded between jellyfish and seawater by using the Thermo Nicolet Avatar 360 Fourier-transform IR method. In the IR-spectrum of jellyfish are observed the more pronouncedly expressed local maximums, detected by Thermo Nicolet Avatar 360 Fourier-transform IR and the DNES method. The measurements demonstrate that two common local maximums are observed in IR-spectra of jellyfish and seawater. These maximums are not observed in the IR-spectrum of the cactus juice and mineral water from Rupite (Bulgaria). Jellyfish contains approximately 97 (w/w) % of water and is more unstable living organism compared to those ones formed the stromatolites. The explanation for this is the smaller concentration of salts in jellyfish and, therefore, the smaller number of local maximums in the IR-spectrum in relation to seawater.

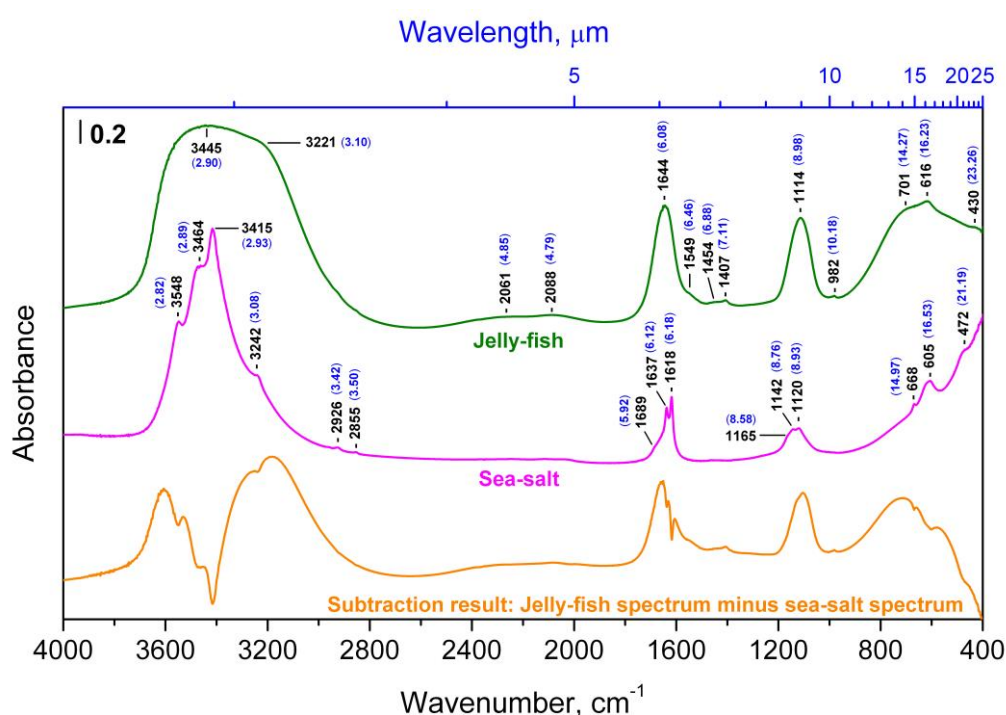
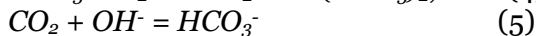
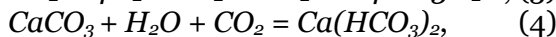


Figure 3. IR-spectrum of seawater obtained from Varna (Bulgaria) and jellyfish *Cotylorhiza tuberculata*, Chalkida (Greece), Aegean Sea

Such a character of IR- and DNES-spectra and distribution of local maximums may prove that hot mineral alkaline water is preferable for origin and maintenance of life compared to other types of water analyzed by these methods. Thus, in hot mineral waters the local maximums in the IR-spectrum are more manifested compared to the local maximums obtained in IR-spectrum of the same water at a lower temperature. The difference in the local maximums from  $t = +20 \text{ }^\circ\text{C}$  to  $t = +95 \text{ }^\circ\text{C}$  at each  $+5 \text{ }^\circ\text{C}$  according to the Student  $t$ -criterion makes up  $p < 0,05$ . These data indicate that the origination of life and living matter depends on the structure and physical chemical properties of water, as well as its temperature and the pH value. The most closed to the IR- and the DNES-spectrum of water, which contains bicarbonates and calcium ions typical for the formation of stromatolites, is the IR-spectrum of the cactus juice. For this reason cactus juice was applied as a model system. The most closed to local maximums in the IR-spectrum of the cactus juice are local maximums in IR-spectra of alkaline mineral water interacting with  $\text{CaCO}_3$  and then seawater.

In connection with these data the following reactions participating with  $\text{CaCO}_3$  in aqueous solutions are important:



The equation (3) shows how some chemosynthetic bacteria use energy from the oxidation of  $\text{H}_2\text{S}$  and  $\text{CO}_2$  to  $\text{S}$  and formaldehyde ( $\text{CH}_2\text{O}$ ). The equation (4) is related to one of the most common processes in nature: in the presence of  $\text{H}_2\text{O}$  and  $\text{CO}_2$ ,  $\text{CaCO}_3$  transforms into  $\text{Ca}(\text{HCO}_3)_2$ . In the presence of hydroxyl  $\text{OH}^-$  ions,  $\text{CO}_2$  transforms into  $\text{HCO}_3^-$  (equation (5)). Equation (6) is valid for the process of formation of the stromatolites – the dolomite layered accretionary structures formed in shallow seawater by colonies of cyanobacteria. In 2010 D. Ward described fossilized stromatolites in the Glacier National Park (USA) [13]. The stromatolites aged 3,5 billion years had lived in warm and hot water in zones of volcanic activity, which could be heated by magma. This suggests that the first living forms evidently evolved in hot geysers [14]. It is known that water in geysers is rich in carbonates, while the temperature is ranged from +100 °C to +150 °C. In 2011 a team of Japanese scientists under the leadership of T. Sugawara showed that life originated in warm or, more likely, in hot water [15]. From an aqueous solution of organic molecules, DNA and synthetic enzymes were created proto cells. Under experimental conditions the initial solution was heated to a temperature close to the water's boiling point +95 °C. Then the temperature was lowered to +65 °C with formation of proto cells with primitive membrane. These experiments are excellent confirmation of the possibility that first organic forms of life originated in hot water.

The above-mentioned data can predict a possible transition from synthesis of small organic molecules under the temperatures +70–100 °C to more complex organic molecules as proteins. There are reactions of condensation-dehydration of amino acids into separate blocks of peptides that occur under alkaline conditions, with  $\text{pH} = 9\text{--}11$ . The important factor in reaction of condensation of two amino acid molecules into dipeptide is allocation of  $\text{H}_2\text{O}$  molecule when a peptide chain is formed. Because the reaction of polycondensation of amino acids is accompanied by dehydration, the  $\text{H}_2\text{O}$  removal from the reaction mixture speeds up the reaction rates. This testifies that formation of early organic forms may have occurred nearby active volcanoes, because at early periods of geological history volcanic activity occurred more actively than during subsequent geological times. However, the dehydration accompanies not only amino acid polymerization, but also association of other small blocks into larger organic molecules, and also polymerization of nucleotides into nucleic acids. Such association is connected with the reaction of condensation, at which from one block a proton is removed, and from another – a hydroxyl group with the formation of  $\text{H}_2\text{O}$  molecule.

In 1969 the possibility of existence of condensation-dehydration reactions under conditions of primary hydrosphere was proven by M. Calvin [16]. From most chemical substances hydrocyanic acid ( $\text{HCN}$ ) and its derivatives – cyanoamid ( $\text{CH}_2\text{N}_2$ ) and dicyanoamid ( $\text{HN}(\text{CN})_2$ ) possess the dehydration ability and the ability to catalyze the process of linkage of  $\text{H}_2\text{O}$  from primary hydrosphere [17]. The presence of  $\text{HCN}$  in primary hydrosphere was proven by S. Miller's early experiments [18]. The chemical reactions with  $\text{HCN}$  and its derivatives are complex with a chemical point of view; in the presence of  $\text{HCN}$ ,  $\text{CH}_2\text{N}_2$  and  $\text{HN}(\text{CN})_2$  the condensation of separate blocks of amino acids accompanied by the dehydration, can proceed at normal temperatures in strongly diluted  $\text{H}_2\text{O}$ -solutions. These reactions show the results of synthesis from separate smaller molecules to larger organic molecules of polymers, e.g. proteins, polycarboxydrates, lipids and ribonucleic acids. Furthermore, the polycondensation reactions catalyzed by  $\text{HCN}$  and its derivatives depend on acidity of water solutions in which they proceed [19]. In acid aqueous solutions with  $\text{pH} = 4\text{--}6$  these reactions do not occur, whereas alkaline conditions with  $\text{pH} = 9\text{--}10$  promote their course. There has not been unequivocal opinion, whether primary water was alkaline, but it is probable that such the  $\text{pH}$  value possessed mineral waters adjoining with basalts, i.e. these reactions could occur at the contact of water with basalt rocks, that testifies our hypothesis.

It should be noted, that the geothermal sources might be used for synthesis of various organic molecules. Thus, amino acids were detected in solutions of formaldehyde  $\text{CH}_2\text{O}$  with

hydroxylamine  $\text{NH}_2\text{OH}$ , formaldehyde with hydrazine ( $\text{N}_2\text{H}_4$ ) in water solutions with HCN, after the heating of a reactionary mixture to  $+95\text{ }^\circ\text{C}$  [20]. In model experiments the reaction products were polymerized into peptide chains that are the important stage towards inorganic synthesis of protein. In a reactionary mixture with a HCN– $\text{NH}_3$  solution in water were formed purines and pyrimidines. In other experiments amino acid mixtures were subjected to influence of temperatures from  $+60\text{ }^\circ\text{C}$  up to  $+170\text{ }^\circ\text{C}$  with formation of short protein-like molecules resembling early evolutionary forms of proteins subsequently designated as thermal proteinoids. They consisted of 18 amino acids usually occurring in protein hydrolyzates. The synthesized proteinoids are similar to natural proteins on a number of other important properties, e. g. on linkage by nucleobases and ability to cause the reactions similar to those catalyzed by enzymes in living organisms as decarboxylation, amination, deamination, and oxidoreduction. Proteinoids are capable to catalytically decompose glucose [21] and to have an effect similar to the action of  $\alpha$ -melanocyte-stimulating hormone [22].

Under certain conditions (temperature, pH) in hot mixture of thermal proteinoids in water solutions are formed elementary structures like proteinoid microspheres with diameter  $5\text{--}10\text{ }\mu\text{m}$  [23]. The best results on polycondensation were achieved with the mixes of amino acids containing aspartic and glutamic acids, which are essential amino acids occurring in all modern living organisms. By morphological features the proteinoid microspheres with a diameter  $\sim 5\text{--}10\text{ }\mu\text{m}$  resemble cell membrane, which in certain conditions ( $\text{pH} = 4\text{--}5$ ) may be double. The catalyst for their formation could serve sulfur and its derivatives which were found in ancient rocks in the form of grains of sulfides, as well as pyrite sands. Synthesis of proteinoid microspheres from a mixture of thermal proteinoids is important because it provides material for the next stage of the evolution of life. This is the stage from disparate organic molecules to organized proteinoid molecules having organized structure and separated from the surrounding environment by the primitive membrane.

In further experiments was used the gas electric discharge analogous to S. Miller's experiments [24]. The first experiments on the modeling of non-equilibrium conditions with gas electric discharge simulating primary atmosphere and electrosynthesis of organic substances from anorganic ones under the energy of the electric field in a primary oxygen-free atmosphere were held in 1953 by S. Miller (USA) [24]. For this aim the mixture of water and gases consisted of hydrogen ( $\text{H}_2$ ), methane ( $\text{CH}_4$ ), ammonia ( $\text{NH}_3$ ) and carbon monoxide ( $\text{CO}$ ) was placed into a closed glass container being exposed by the pulsating electrical spark discharges at the temperature of boiling water; oxygen was not allowed into the unit. After the processing the reaction mixture by the electric discharge it was found that from the initial inorganic substance was synthesized the organic compounds – aldehydes and amino acids. Experiments found that approximately  $\sim 10\text{--}15\%$  of carbon was transferred into an organic form. However, about  $\sim 2\%$  of carbon was detected in the amino acids, the most common of which was glycine. The initial analysis showed the presence in the reaction mixture obtained after the processing by spark electric discharge 5 amino acids. A more complete analysis carried out in 2008 [25], showed the formation by electrosynthesis in the reaction mixture 22 amino acids having from 5 to 20 carbon atoms in the molecule. Interestingly is that along with the amino acids in the reaction mixture after the treatment with electric spark discharges were detected trace amounts of nucleic acid precursors – nucleosides.

It should be noted that in the implementation of the gas discharge effect as well as in experiments of S. Miller are modeled extreme non-equilibrium conditions with gas electric discharge, resulting that in a thin layer of air gap with thickness  $\sim 100\text{ }\mu\text{m}$  are formed reactive radicals reacting with each other to form new compounds (electrosynthesis). Such extreme conditions are thought to have occurred in the primary oxygen-free atmosphere of the Earth, which supposedly consisted of a mixture of water and gases –  $\text{H}_2$ ,  $\text{CH}_4$ ,  $\text{NH}_3$  and  $\text{CO}$ , subjected to spark electrical discharges (lightning) under the conditions of high solar (UV) and geothermal activity.

The analogous experiment was conducted by the authors under laboratory conditions. According to our previous experiments, the first living structures may have evolved in warm and hot mineral water with a high content of bicarbonate ( $\text{HCO}_3^-$ ) anions, cations of alkali metals ( $\text{Na}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ) and deuterium in the form of HDO [26]. There occurred gas electric discharge (lightning) in the primordial atmosphere close to the water surface. In the course of experiment was used the similar gas electric discharge on water drops placed on the electrode of the device for gas electric discharge formation. The similar composition and water temperature were modeled on



the electrode of the gas discharge device made of hostafan, with electric voltage – 15 kV, electric impulse duration – 10  $\mu$ s; electric current frequency – 15 kHz, wherein the air gap layer on the boundary with water sample was formed the electrical discharge, similar to plasma phenomena (lightning) and the electrostatic discharge on the surface of organic and inorganic samples of various kinds. Water drops were heated up to the boiling point in an electric field of high frequency and the electric discharge was applied, analogous as that in the primordial atmosphere. As a result, an organized structure with a size of 1,2–1,4 mm was formed in interelectrode space (Fig. 4). It was formed as a result of accretion of smaller elementary structures sized up 5–10  $\mu$ m into the biggest structure having the size 1,2–1,4 mm and concentrated in the space where the electric field is applied.

It should be noted that no structure was organized in a control sample of water being placed on the electrode. Before its placement on the electrode, the water was heated to boiling point and then cooled down. The structure organization increased with the increase of the duration of the gas electric discharge. Moreover, in experiments was observed the formation of small structures and their further “adjoining” to the larger structure. The large structure was preserved the original size for some time in the absence of the electric discharge.



Figure 4. The organized structure in water sample subjected to the temperature +100 °C in the electric field of high voltage and frequency (I. Ignatov, 2014). The material of the electrode – hostafan; the electric voltage – 15 kV, the electric impulse duration – 10  $\mu$ s; the electric current frequency – 15 kHz.

This experiment shows that self-organization in water under certain external thermal conditions may take place. In natural conditions water is heated up to +100 °C by the magma. The structure formed from heated water was evidently a result of self-organization. Living organisms are complex self-organizing systems. Thermodynamically they belong to the open systems because they constantly exchange substances and energy with the environment. The changes in the open systems are relatively stable in time. The stable correlation between components in an open system is called a dissipative structure. According to I. Prigozhin, the formation of dissipative structures and the elaboration to living cells is related to changes in entropy [27].

Taking into account these views it may be concluded that the initial stage of evolution, apparently, was connected with formation at high temperature of the mixtures of amino acids and nitrogenous substances – analogues of nucleic acids. Such synthesis is possible in aqueous

solutions under thermal conditions in the presence of  $\text{H}_3\text{PO}_4$ . The next stage is the polycondensation of amino acids into thermal proteinoids at temperatures 65–95 °C. After that in a mix of thermal proteinoids in hot water solutions were formed membrane like structures.

Our data are confirmed by experiments of T. Sugawara (Japan), who in 2011 created the membrane like proto cells from an aqueous solution of organic molecules, DNA and synthetic enzymes under the temperature close to the water boiling point +95 °C [28]. This data confirm the possibility that first organic forms of life might have been originated in hot mineral water.

### Physical-chemical Properties of $\text{D}_2\text{O}$ -solutions

The numerous studies carried out by us with various biological objects in  $\text{D}_2\text{O}$ , proved that when biological objects being exposed to water with different deuterium content, their reaction varies depending on the isotopic composition of water (the content of deuterium in water) and magnitude of isotope effects determined by the difference of constants of chemical reactions rates  $k_{\text{H}}/k_{\text{D}}$  in  $\text{H}_2\text{O}$  and  $\text{D}_2\text{O}$ . The maximum kinetic isotopic effect observed at ordinary temperatures in chemical reactions leading to rupture of bonds involving hydrogen and deuterium atoms lies in the range  $k_{\text{H}}/k_{\text{D}} = 5\text{--}8$  for C–H versus C–D, N–D versus N–D, and O–D versus bonds [29].

The chemical structure of  $\text{D}_2\text{O}$  molecule is analogous to that one for  $\text{H}_2\text{O}$ , with small differences in the length of the covalent H–O-bonds and the angles between them. The difference in nuclear masses stipulates the isotopic effects, which may be sufficiently essential for the  $^1\text{H}/\text{D}$  pair. As a result, physical-chemical properties of  $\text{D}_2\text{O}$  differ from  $\text{H}_2\text{O}$ :  $\text{D}_2\text{O}$  boils at  $t = +101,44$  °C, freezes at  $t = +3,82$  °C, has maximal density at  $t = +11,2$  °C ( $1,106$  g/cm<sup>3</sup>) [30]. In mixtures of  $^2\text{H}_2\text{O}$  with  $\text{H}_2\text{O}$  the isotopic exchange occurs with high speed with the formation of semi-heavy water ( $^1\text{HDO}$ ):  $\text{D}_2\text{O} + \text{H}_2\text{O} = ^1\text{HDO}$ . For this reason deuterium presents in smaller content in aqueous solutions in form of  $^1\text{HDO}$ , while in the higher content – in form of  $\text{D}_2\text{O}$ . The chemical reactions in  $\text{D}_2\text{O}$  are somehow slower compared to  $\text{H}_2\text{O}$ .  $\text{D}_2\text{O}$  is less ionized, the dissociation constant of  $\text{D}_2\text{O}$  is smaller, and the solubility of the organic and inorganic substances in  $\text{D}_2\text{O}$  is smaller compared to these ones in  $\text{H}_2\text{O}$ . Due to isotopic effects the hydrogen bonds with the participation of deuterium are slightly stronger than those ones formed of hydrogen.

The comparative analysis of IR-spectra of  $\text{H}_2\text{O}$  solutions and its deuterated analogues ( $\text{D}_2\text{O}$ , HDO) is of considerable interest for biophysical studies, because at changing of the atomic mass of hydrogen by deuterium atoms in  $\text{H}_2\text{O}$  molecule their interaction will also change, although the electronic structure of the molecule and its ability to form H-bonds, however, remains the same. The IR spectra of water usually contain three absorption bands, which can be identified as 1 – absorption band of the stretching vibration of OH<sup>-</sup> group; 2 – absorption band of the first overtone of the bending vibration of the molecule HDO; 3 – absorption band of stretching vibration of OD<sup>-</sup> group. OH<sup>-</sup> group is able to absorb much infrared radiation in the infrared region of the IR-spectrum. Because of its polarity, these groups typically react with each other or with other polar groups to form intra- and intermolecular hydrogen bonds. The hydroxyl groups not involved in formation of hydrogen bonds are usually given the narrow bands in IR spectrum and the associated groups – broad intense absorption bands at lower frequencies. The magnitude of the frequency shift is determined by the strength of the hydrogen bond. Complication of the IR spectrum in the area of OH<sup>-</sup> stretching vibrations can be explained by the existence of different types of associations, a manifestation of overtones and combination frequencies of OH<sup>-</sup> groups in hydrogen bonding, as well as the proton tunneling effect (on the relay mechanism. Such complexity makes it difficult to interpret the IR spectrum and partly explains the discrepancy in the literature available on this subject.

The local maximums in IR-spectra reflect vibrational-rotational transitions in the ground electronic state; the substitution with deuterium changes the vibrational-rotational transitions in  $\text{H}_2\text{O}$  molecule that is why it appears other local maximums in IR-spectra. In the water vapor state, the vibrations involve combinations of symmetric stretch ( $\nu_1$ ), asymmetric stretch ( $\nu_3$ ) and bending ( $\nu_2$ ) of the covalent bonds with absorption intensity ( $\text{H}_2\text{O}$ )  $\nu_1; \nu_2; \nu_3 = 2671; 1178,4; 2787,7$  cm<sup>-1</sup>. For liquid water the absorption bands are observed in other regions of the IR-spectrum, the most intense of which are located at 2100 cm<sup>-1</sup> and 710–645 cm<sup>-1</sup>. For  $\text{D}_2\text{O}$  molecule these ratio compiles 2723,7; 1403,5 and 3707,5 cm<sup>-1</sup>, while for HDO molecule – 2671,6; 1178,4 and 2787,7 cm<sup>-1</sup>. HDO (50 mole%  $\text{H}_2\text{O}$  + 50 mole%  $^2\text{H}_2\text{O}$ ; ~50 % HDO, ~25 %  $\text{H}_2\text{O}$ , ~25 %  $\text{D}_2\text{O}$ ) has local maxima in IR-

spectra at 3415  $\text{cm}^{-1}$ , 2495  $\text{cm}^{-1}$  1850  $\text{cm}^{-1}$  and 1450  $\text{cm}^{-1}$  assigned to OH<sup>-</sup>-stretch, OD<sup>-</sup>-stretch, as well as combination of bending and libration and HDO bending respectively.

In the IR-spectrum of liquid water absorbance band considerably broadened and shifted relative to the corresponding bands in the spectrum of water vapor. Their position depends on the temperature [31]. The temperature dependence of individual spectral bands of liquid water is very complex. Furthermore, the complexity of the IR-spectrum in the area of OH<sup>-</sup> stretching vibration can be explained by the existence of different types of H<sub>2</sub>O associations, manifestation of overtones and composite frequencies of OH<sup>-</sup> groups in the hydrogen bonds, and the tunneling effect of the proton (for relay mechanism). Such complexity makes it difficult to interpret the spectrum and partly explains the discrepancy in the literature available on this subject.

In liquid water and ice the IR-spectra are far more complex than those ones of the vapor due to the vibrational overtones and combinations with librations (restricted rotations, i.g. rocking motions). These librations are due to the restrictions imposed by hydrogen bonding (minor L<sub>1</sub> band at 395,5  $\text{cm}^{-1}$ ; major L<sub>2</sub> band at 686,3  $\text{cm}^{-1}$ ; for liquid water at 0 °C, the absorbance of L<sub>1</sub> increasing with increasing temperature, while L<sub>2</sub> absorbance decreases but broadens with reduced wave number with increasing temperature [32]. The IR spectra of liquid water usually contain three absorbance bands, which can be identified on absorption band of the stretching vibration of OH<sup>-</sup> group; an absorption band of the first overtone of the bending vibration of the molecule HDO and absorption band of stretching vibration of OD<sup>-</sup> group. Hydroxyl group OH<sup>-</sup> is able to absorb much infrared radiation in the infrared region of the IR-spectrum. Because of its polarity, these groups typically react with each other or with other polar groups to form intra- and inter-molecular hydrogen bonds. The hydroxyl groups, which are not involved in formation of hydrogen bonds, usually produce the narrow bands in IR spectrum, while the associated groups – broad intense absorbance bands at lower frequencies. The magnitude of the frequency shift is determined by the strength of the hydrogen bond. The complication of the IR spectrum in the area of OH<sup>-</sup> stretching vibrations can be explained by the existence of different types of associations of H<sub>2</sub>O molecules, a manifestation of overtones and combination frequencies of OH<sup>-</sup> groups in hydrogen bonding, as well as the proton tunneling effect (on the relay mechanism).

The assignment of main absorption bands in the IR-spectrum of liquid water is given in Table 2. The IR spectrum of H<sub>2</sub>O molecule was examined in detail from the microwave till the middle (4–17500  $\text{cm}^{-1}$ ) visible region and the ultraviolet region – from 200 nm<sup>-1</sup> to ionization limit at 98 nm<sup>-1</sup> [33]. In the middle visible region at 4–7500  $\text{cm}^{-1}$  are located the rotational spectrum and the bands corresponding to the vibrational-rotational transitions in the ground electronic state. In the ultraviolet region (from 200 nm<sup>-1</sup> to 98 nm<sup>-1</sup>) are located the bands corresponding to transitions from the excited electronic states close to the ionization limit in the electronic ground state. The intermediate region of the IR-spectrum – from 570 nm to 200 nm corresponds to transitions to higher vibrational levels of the ground electronic state.

Table 2: The assignment of main frequencies in IR-spectra of liquid water H<sub>2</sub>O and D<sub>2</sub>O

<b>Main vibrations of liquid H<sub>2</sub>O and <sup>2</sup>H<sub>2</sub>O</b>				
<b>Vibration(s)</b>	<b>H<sub>2</sub>O (t = +25 °C)</b>		<b>D<sub>2</sub>O (t = +25 °C)</b>	
	<b>v, cm<sup>-1</sup></b>	<b>E<sub>0</sub>, M<sup>-1</sup> cm<sup>-1</sup></b>	<b>v, cm<sup>-1</sup></b>	<b>E<sub>0</sub>, M<sup>-1</sup> cm<sup>-1</sup></b>
Spinning v <sub>1</sub> + deformation v <sub>2</sub>	780–1645	21,65	1210	17,10
Composite v <sub>1</sub> + v <sub>2</sub>	2150	3,46	1555	1,88
Valence symmetrical v <sub>1</sub> , valence asymmetrical v <sub>3</sub> , and overtone 2v <sub>2</sub>	3290–3450	100,65	2510	69,70

The results of IR-spectroscopy with Infra Spec VFA-IR device show that at  $\lambda = 4,1 \mu\text{m}$ , even at low concentrations of deuterium of 0,35 and 0,71%, there is observed a decline in the local maximums relative to the local maximum of 100% pure water (the local maximums in IR-spectra reflect vibrational-rotational transitions in the ground electronic state because at changing the

atomic mass of hydrogen and deuterium atoms in the water molecule their interaction will also change, although the electronic structure of the molecule and its ability to form H-bonds, however, remains the same; with the substitution with deuterium the vibrational-rotational transitions are changed, that is why it appears other local maximums in IR-spectra. The result is reliable regarding the content of deuterium in natural waters from 0,015–0,03%.

At further transition from H<sub>2</sub>O monomers to H<sub>4</sub>O<sub>2</sub> dimer and H<sub>6</sub>O<sub>3</sub> trimer absorption maximum of valent stretching vibrations of the O–H bond is shifted toward lower frequencies ( $\nu_3 = 3490 \text{ cm}^{-1}$  and  $\nu_1 = 3280 \text{ cm}^{-1}$ ) [34] and the bending frequency increased ( $\nu_2 = 1644 \text{ cm}^{-1}$ ) because of the hydrogen bonding. The increased strength of hydrogen bonding typically shifts the stretch vibration to lower frequencies (red-shift) with greatly increased intensity in the infrared due to the increased dipoles. In contrast, for the deformation vibrations of the H–O–H, it is observed a shift towards higher frequencies. The absorption bands at 3546 and 3691  $\text{cm}^{-1}$  were attributed to the stretching modes of the dimer [(H<sub>2</sub>O)<sub>2</sub>]. These frequencies are significantly lower than the valence modes of  $\nu_1$  and  $\nu_3$  vibrations of isolated H<sub>2</sub>O molecules at 3657 and 3756  $\text{cm}^{-1}$  respectively). The absorption band at 3250  $\text{cm}^{-1}$  represents overtones of deformation vibrations. Among the frequencies between 3250 and 3420  $\text{cm}^{-1}$  it is possible the Fermi resonance (this resonance is a single substitution of intensity of one fluctuation by another fluctuation when they accidentally overlap each other). The absorption band at 1620  $\text{cm}^{-1}$  is attributed to the deformation mode of the dimer. This frequency is slightly higher than the deformation mode of the isolated H<sub>2</sub>O molecule (1596  $\text{cm}^{-1}$ ). A shift of the band of deformation vibration of water in the direction of high frequencies at the transition from a liquid to a solid state is attributed by the appearance of additional force, preventing O–H bond bending. The deformation absorption band in IR-spectrum of water has a frequency at 1645  $\text{cm}^{-1}$  and very weak temperature dependence. It changes little in the transition to the individual H<sub>2</sub>O molecule at a frequency of 1595  $\text{cm}^{-1}$ . This frequency is found to be sufficiently stable, while all other frequencies are greatly affected by temperature changes, the dissolution of the salts and phase transitions. It is believed that the persistence of deformation oscillations is stipulated by processes of intermolecular interactions, e.g. by the change in bond angle as a result of interaction of H<sub>2</sub>O molecules with each other, as well as with cations and anions.

Thus, the study of the characteristics of the IR spectrum of D<sub>2</sub>O-solutions allows to answer the question not only on the physical parameters of the molecule and the covalent bonds at isotopic substitution with deuterium, but also to make a certain conclusion on associative environment in water. The latter fact is important in the study of structural and functional properties of water associates and its isotopomers at the isotopic substitution with deuterium. The substitution of H with D affects the stability and geometry of hydrogen bonds in an apparently rather complex way and may through the changes in the hydrogen bond zero-point vibration energies, alter the conformational dynamics of hydrogen (deuterium)-bonded structures of macromolecules as DNA and proteins in D<sub>2</sub>O.

### Conclusions

The experimental data testify that the origination of first organic forms of life depends on physical-chemical properties of water and external factors – temperature, pH, electric discharges and isotopic composition. Hot mineral alkaline water interacting with CaCO<sub>3</sub> is most closed to these conditions. Next in line with regard to its quality is seawater. For chemical reaction of dehydration-condensation to occur in hot mineral water, water is required to be alkaline with the pH ranged 9–11. In warm and hot mineral waters the local maximums in IR-spectra from 8 to 14  $\mu\text{m}$  were more expressed in comparison with the local maximums measured in the same water samples with lower temperature. The content of deuterium in hot mineral water may be increased due to the physical chemical processes of the deuterium accumulation stipulated by the solar radiation, volcanic geothermal processes and electric discharges in the atmosphere. These natural processes could lead to the enrichment of the hydrosphere by deuterium in the form of HDO which evaporates more slowly than H<sub>2</sub>O, but condenses faster. If the primary hydrosphere really contained HDO, that this may explain the thermal stability of the first organic life forms in the hot mineral water, as the thermal stability of deuterated macromolecules like DNA and proteins in D<sub>2</sub>O solutions is somewhat higher than their protonated forms due to isotopic effects of deuterium.

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### **Могли ли первые органические формы жизни возникнуть в горячей минеральной воде с HDO?**

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**Аннотация.** Изучен состав воды, температура и значение рН в экспериментах по моделированию первичной гидросферы и возможных условий возникновения органических форм жизни в горячей минеральной воде с НДО. Для этой цели проведены исследования горячей минеральной и морской воды из Болгарских источников методами ИК- и ДНЭС-спектроскопии. В качестве модельных систем использовали сок кактуса *Echinopsis pachanoi* и средиземноморскую медузу *Cotylorhiza tuberculata*. Рассмотрены реакции конденсации-дегидратации в щелочных водных растворах со значением рН = 9–10, результатом которых является синтез из мелких молекул более крупных органических молекул полимеров как короткие полипептиды, а также механизмы аккумуляции дейтерия в форме НДО в горячей минеральной воде. Показано, что горячие минеральные воды с температурой от +65 °С до +95 °С и значением рН от 9 до 11 более пригодны для возникновения жизни, чем другие исследованные образцы воды. В горячей минеральной воде локальные максимумы в ИК-спектре проявлялись больше всего, по сравнению с локальными максимумами в ИК-спектре той же воды при более низкой температуре. Разница в локальных максимумах от +20 °С до +95 °С при увеличении температуры на каждый +5 °С составила в соответствии с *t*-критерием Стьюдента  $p < 0,05$ .

**Ключевые слова:** дейтерий, гидросфера, эволюция, возникновение жизни, ИК-спектроскопия, ДНЭС.