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Influence of Pr³⁺ ion doped in Infra-red spectroscopy and Elastic properties Co ferrites

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ABSTRACT

Pr³⁺ doped $Pr_xCoFe_{2-x}O_4$ (x = 0.0, 0.025,0.05,0.075, 0.1)ferrite nanoparticles prepared by using sol-gel auto-combustion method. The prepared samples sintered at 600 ⁰C for 4 hours. Infrared spectra carried out at room temperature in the wavenumber range of 300–800 cm⁻¹. The IR spectra show two major absorption bands. High frequency bands 'v₁' is assigned to the tetrahedral and low frequency bands 'v₂' is assigned to the octahedral sites of complex. Force constant for the tetrahedral and octahedral site was determined by using IR data. The values of Force constant use to calculate the Stiffness constants (C_{11} and C_{12}). Using the values of stiffness constants; Elastic moduli such as Young's modulus, Rigidity modulus, Poisson's ratio and Debye temperature are calculated.

Keywords: Sol-gel method, Ferrites, infra- red spectroscopy, elastic property.

INTRODUCTION

Ferrites have been the subject of extensive investigation in last few decades due to their wide range of applications such as information storage systems, gas sensors, microwave devices and magnetic recording and electronic industries [1]. The possibility of mixing metals in different compositions makes ferrites very attractive. Spinel ferrites with cubic structure are one of the most attractive materials for magnetic, catalysis,

microelectronic and gas sensing applications [2-5]. Cobalt ferrites have been regarded for high density magnetic recording media because of its moderate saturation magnetization, high coercivity, mechanical hardness and chemical stability [6]. Ferrites samples prepared several methods. Such as Solid state reaction [7], Hydrothermal preparation [8], Micro emulsion method [9], Oxalate Precursors method [10]. Now a day, among these various preparation methods of the ferrite nanoparticles, the Sol–gel auto-combustion method commonly used. The elastic constant is important because of them elucidates the nature of the binding forces in the solid. The elastic modulus represents the mechanical strength fracture toughness and thermal stock resistance [11]. $Pr_xCoFe_{2-x}O_4$ (x = 0.0, 0.025, 0.05, 0.075,0.1) were synthesized by Sol-gel Auto-Combustion method and also examine the effect of Pr^{3+} substituted in Co ferrites on the elastic properties.

METHODOLOGY

Nano crystalline ferrite powders with compositions of $Pr_xCoFe_{2-x}O_4$ (x =0.0, 0.025,0.05, 0.075, 0.1) were synthesized by Sol-gel Auto-Combustion technique. The metal nitrates (A.R. grade with 99.8 % purity) Cobalt nitrate. Praseodymium nitrate $Pr(NO_3)_3.6H_2O$, Cobalt nitrate $Co(NO_3)_2$, $6(H_2O)$ Ferric Nitrate $(Fe(NO_3)_3 \cdot 9H_2O)$, and Citric acid $(C_6H_8O_7 \cdot H_2O)$ were used as starting materials. Reaction was carried out in air atmosphere without protection of inert gases.

The molar ratio of metal nitrates to citric acid was taken as 1:3. The metal nitrates were dissolved together in a minimum amount of double distilled water to get a clear solution. An aqueous solution of citric acid was mixed with metal nitrates solution, then ammonia solution was slowly added to adjust the pH ≅ 7. The mixed solution was kept on to a hot plate with continuous stirring at 90 °C. During evaporation, the solution became viscous and finally formed a viscous brown gel. When finally all water molecules were removed from the mixture, viscous gel began frothing. After few minutes, the gel automatically ignited and burnt with glowing flints. The decomposition reaction

would not stop before the whole citrate complex was consumed. The auto-combustion was completed within a minute, yielding the brown-coloured ashes termed as a precursor. The as prepared powder then annealed at 600 °C for 4 hrs.

The prepared samples were characterized by X-ray investigations. Part of the powder was X-ray examined by Phillips X-ray diffractometer (Model 3710) using Cu-K$_\alpha$ radiation (λ=1.5405Å). The scanning step was 0.02° and scanning rate was 2 ° /min. The infrared spectra of all the samples were recorded at room temperature in the range 300 cm^{-1} to 800 cm^{-1}.

RESULTS AND DISCUSSIONS

X-ray analysis:
It is obvious from the XRD patterns (Fig. 1) of $Pr_xCoFe_{2-x}O_4$ that the patterns were indexed and cubic lattice was observed. The results indicate that the material has a well-defined crystalline single phase belonging to the fcc system. The diffraction pattern gave a verification of the presence of spinel crystal structure for each composition of $Pr_xCoFe_{2-x}O_4$.Since the XRD patterns of $PrFeO_3$ phase for x < 0.075, we can draw a conclusion thatPr^{3+}has entered the lattice of ferrite successfully.[12].

It is observed that the cubic spinel phase coexists with some amount of $PrFeO_3$phase in x ≥ 0.075.As $CoFe_2O_4$ is nearly inverse spinel, this anomaly in the intensity distribution may arise from the following facts: (i) differences in the cation distribution among the [A] and [B] sites in the spinel lattice and; (ii) phase inhomogeneity, in particular presence of ferrite phases with different cation distributions. The ionic radius of Pr^{3+} ions is 1.13Å which is larger than that of Fe^{3+} (0.67 Å), and the amount of Fe^{3+}ions substituted by Pr^{3+}ions is limited, thus redundant Pr^{3+}ions aggregates on the grain boundaries forming $PrFeO_3$ phase. As a result, there was a limit for the replacement of Fe^{3+} with Pr^{3+}, and the maximum was x = 0.05 in our experiment .The values of lattice constant and x-ray density calculated are given by Table.1
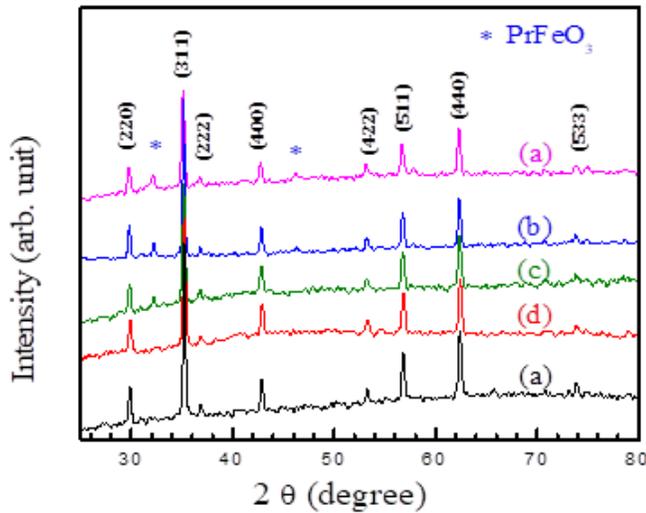
Fig. 1: X-Ray diffraction patterns of (a) x = 0.0, (b) x = 0.025, (c) x = 0.050, (d) x = 0.075 and (e) x = 0.1 for $Pr_xCoFe_{2-x}O_4$

The lattice constant 'a' increases from 8.373- 8.423 Å with increase in Pr^{3+} substitution. The ionic radius of Pr^{3+} and Fe^{3+} ions is 1.13 and 0.67 respectively, therefore, when Fe^{3+} was replacedby$Pr^{3+}$the lattice constant should aggrandize as the content of the Pr^{3+}increased.The relationship between Pr^{3+} substitution and lattice parameter can be explained by Vegard's law [13].

Infra-red Spectroscopy:
The obtained IR spectrum in the range 200-800 cm^{-1} for nanocrysatlline $Pr_xCoFe_{2-x}O_4$ ferrite system is illustrated in Fig. 2. The observed variation in absorption bands for the present investigated system is listed in Table 1. It is observed from Table 1. IR spectrum of the synthesized samples displayed two absorption bands characteristic of the spinel ferrites at 707-757 and 454-477 cm^{-1}. The higher frequency band (v_1) was due to stretching vibration of tetrahedral metal oxygen bond and the lower frequency band (v_2) was due to octahedral metal oxygen bond [13]. The band positions for all the investigated composition are given in Table 1. The change in the lattice constant is responsible for this shift of the center frequencies.

The increase in the unit cell dimensions due to the replacement of Fe^{3+} ions by larger ionic radius Pr^{3+} ions affects the Fe^{3+}–O^{2-}stretching vibrations and this is a

prominent cause of change in band positions. The change in the frequency of v_1 stretching band indicates the preference of Pr^{3+} ions to occupy the octahedral sites. The peak intensity of frequency bands slightly changes within increasing Pr^{3+}substitution. It is known that the intensity ratio is a function of the change of dipole moment with the inter-nuclear distance (dμ/dr) [14]. This value represents the contribution of the ionic bond Fe-O in the lattice. Furthermore, it is observed from Fig. 5.10that the normal mode of vibration of tetrahedral cluster (v_1) is higher than that of octahedral cluster(v_2), which is attributed to the shorter bond length of tetrahedral cluster and longer bond length of octahedral cluster. Using the analysis of Waldron [13], the force constant K_0 and K_t were calculated
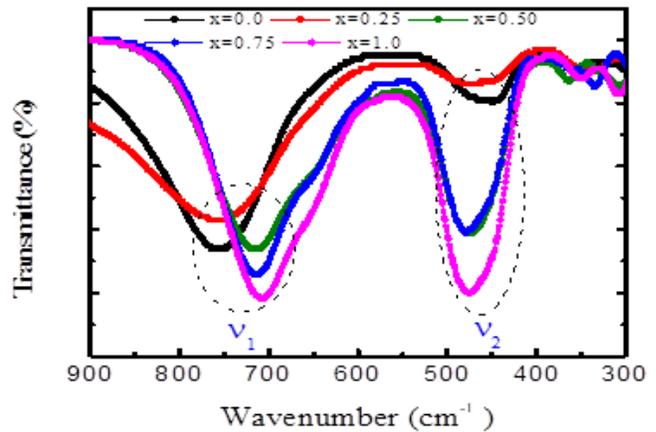
Fig 2: Infrared spectra for the series $Pr_xCoFe_{2-x}O_4$.

According to Waldron the force constant K_t and K_o for respective sites are given by:

$$K_t = 7.62 \times M_1 \times v_1^2 \times 10^{-3} \qquad (1)$$

$$K_0 = 10.62 \times \frac{M_2}{2} \times v_2^2 \times 10^{-3} \qquad (2)$$

Where, K_0 = force constant on octahedral site, K_t = force constant on tetrahedral site, M_1=Molecular weight of tetrahedral site, M_2=Molecular weight of octahedral site, v_1=Corresponding center frequency on tetrahedral site, v_2 = Corresponding center frequency on octahedral site. The values of force are summarized in Table 1. The force constant Kt decreasing with the increasing Pr^{3+}content whereas K_O increases with the increasing in Pr^{3+}. This variation can be related with the difference in ionic radii

of Fe^{3+} and Pr^{3+} ions and their occupancy at A and B sites. Analysis of IR spectra with crystallographic knowledge helps us to determine the Debye temperature and elastic properties. The Debye temperature (θI) of all samples was calculated using the wave number of IR bands.[14].

$$\theta l = hCVac/k$$

Where, h = h/2π, k is Boltzman constant, C is velocity of light (C = 3x10⁸ cm/s) and V_{av} is average wave number of bands. Variation of Debye temperature with Pr content is shown in Fig. 3

Elastic properties:

The elastic properties were determined using infrared spectroscopy [15-16] These elastic moduli were calculated using the values of lattice constant 'a', X-ray density 'dx', pore fraction 'f' and force constant 'K'. Values of lattice constant, X-ray density and pore fraction are listed in Table 1. The average force constant (K) was calculated using the following relation:

$$k = kt + ko/2 \qquad (4)$$

The bulk modulus of term stiffness constant C_{11} was calculated using relation [17] Stiffness constant

$$C_{11} = \frac{K}{a} \qquad (5)$$

Where K is average force constant and a is lattice constant.

$$\text{Stiffness constant} \ (C_{12}) = \frac{\sigma \times C11}{(1-\sigma)} \qquad (6)$$

where σ is Poisson ratio and 'a' is the lattice constant. The Poisson ratio is function of pore fraction (σ = 0.324 x 1 -1.043f). Values of Poisson ratio are presented in Table 2. The Poisson ratio is ranges in between 0.276 and 0.266, these values are lie in the range of -1 to 0.5, which is in conformity with the theory of isotropic elasticity. [16]

Variation of stiffness constants (C11 and C12) as a function of Pr^{3+}content is shown in table.2. It is observed from Table 2; both the stiffness constant was decreases with increase in Pr^{3+} substitution. The values of

Poisson's ratio were calculated using the relation discussed elsewhere [17] and the values are presented in table 2. Use of stiffness constant are affected by two factors; i.e. the tightness of bonding between the atoms and force constant. In present system bonds between Fe^{3+} and Pr^{3+} atoms is residual bond and due to this stiffness constant are decreases with increasing Al content. These two(3)stiffness constants are further used to calculate the various elastic constants such as; Young's modulus (E), bulk modulus (K) and modulus of rigidity (G) [14].The other elastic moduli for cubic structure are calculated using following relation [18.].

$$\text{Rigidity modulus} \ (G) = \frac{E}{2(\sigma+1)} \qquad (7)$$

The rigidity modulus (G) is calculated using the relation 3 and the variation are presented in Table 2. It can be observed from Table 2 the values of rigidity modulus increased with Pr^{3+} substitution. B ,G, E decrease with increase in both Pr^{3+} it indicates that deformation of the solid is easy and the solid has less tendency to spring back to its equilibrium position. the Young's modulus, Bulk modulus and modulus of rigidity decreases with the increasing Pr^{3+} content. The decreases in elastic moduli may be due to the interatomic binding between various atoms in the spinel lattice. [15]. The inter-atomic bonding between the various atoms weakens continuously with the addition of Pr^{3+} content and therefore elastic moduli decreases with the increasing Pr^{3+}content. In Fe^{3+}-Pr^{3+} ferrite repulsion between electrons may be increased with the increasing Pr^{3+} content.[14]

The longitudinal elastic wave velocity (V_L) and transverse (Shear) wave velocity (V_S) was calculated using following equations,

$$\text{Longitudinal velocity} \ V_L = \left(\frac{C_{11}}{\rho}\right)^{1/2} \qquad (8)$$

$$\text{Transverse (Shear) velocity} \ V_S = \left(\frac{G}{\rho}\right)^{1/2} \qquad (9)$$

Where G is rigidity modulus with correct zero pore fraction. The values of V_l and V_s used to calculate mean

wave velocity (V_m) which used to calculate Debye temperature was calculated using formula

Debye temperature

$$\theta_E = \frac{h}{k}\left[\frac{3\rho q N_A}{4\Pi M}\right]^{1/3} \times V_m \qquad (10)$$

Where h is planks constant, k is Boltzmann's constant, M is molecular weight, q is number if atom in the unit formula and V_m mean wave velocity

$$\frac{3}{V_m^3} = \frac{1}{V_l^3} + \frac{2}{V_s^3} \qquad (11)$$

The values of longitudinal wave, shearing wave and mean wave velocity are calculated using relation 4, 5 and 7 respectively. It is observed from Table 3, the longitudinal elastic wave velocity is decreases whereas transverse (Shear) wave velocity increased with Pr^{3+} substitution. The values of wave velocities are similar to other ferrites those obtained from UPT method [20]. The variation of Debye temperature (θ_E) is given Table 3. The Debye temperature increased with Pr^{3+} substitution. It suggested that lattice vibrations are hindered due to Pr^{3+} substitution. This may be due to the fact that strength of interatomic bonding increases with concentration (x and y) as supported by our results on the variation of elastic moduli.

Table 1: Lattice constant (a), X-ray Density (dx), Band Position (v_1 and v_2), force constant (K_o and K_t) of $Pr_xCoFe_{2-x}O_4$

Comp. x	'a' (Å)	'd$_x$' (g/cm³)	Band position		Force constant	
			v_1(cm⁻¹)	v_2 (cm⁻¹)	K_0 x 10^5 (dyne/cm)	K_t x 10^5 (dyne/cm)
0.0	8.3732	5.309	757	454	1.293	2.441
0.025	8.3761	5.352	754	461	1.272	2.399
0.05	8.3835	5.386	718	471	1.266	2.147
0.075	8.3984	5.405	713	477	1.227	2.118
0.1	8.4234	5.404	707	475	1.133	2.121

Table 2: Mean force constant (K), pore fraction Stiffness constant (C_{11} and C_{12}), of **$Pr_xCoFe_{2-x}O_4$**.

Comp. x	Mean Force constant (K) (K_t+K_o)	Pore Fraction	Passion's ratio σ	C_{11}	C_{12}
0.0	1.8671	0.142	0.276	222.98	85.01
0.025	1.8356	0.148	0.274	219.15	82.70
0.05	1.7068	0.154	0.272	203.59	76.05
0.075	1.6725	0.163	0.269	199.14	73.25
0.1	1.6268	0.172	0.266	193.13	69.95

Table 3:. Rigidity modulus (G)Young modulus(Y), Bulk Modulus (B), longitudinal elastic wave velocity (V_L) and transverse (Shear) wave velocity (V_S), mean wave velocity (V_m), Debye temperature (θ_E) of $Pr_xCoFe_{2-x}O_4$.

Comp. x	G	Y	B	V_L	V_S	V_m	θ_E
0.0	91.391	176.05	131.00	6480.77	3280.30	3641.74	498
0.025	91.880	173.83	128.19	6399.03	3288.72	3651.09	500
0.05	87.350	162.23	118.56	6148.18	3277.82	3638.99	499
0.075	88.532	159.75	115.22	6069.94	3280.95	3642.46	501
0.1	89.068	155.94	111.01	5978.18	3295.66	3658.79	503

CONCLUSION

The Pr^{3+} substituted $Pr_xCoFe_{2-x}O_4$ (x=0.0 ,0.025, 0.05, 0.075,0.1) ferrite nanoparticles were prepared by using sol-gel auto-combustion method. The Elastic properties such as elastic wave velocity, elastic constant and Debye temperature calculated using XRD and IR data. The elastic constants decreased with the increasing in Pr^{3+} content.

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